

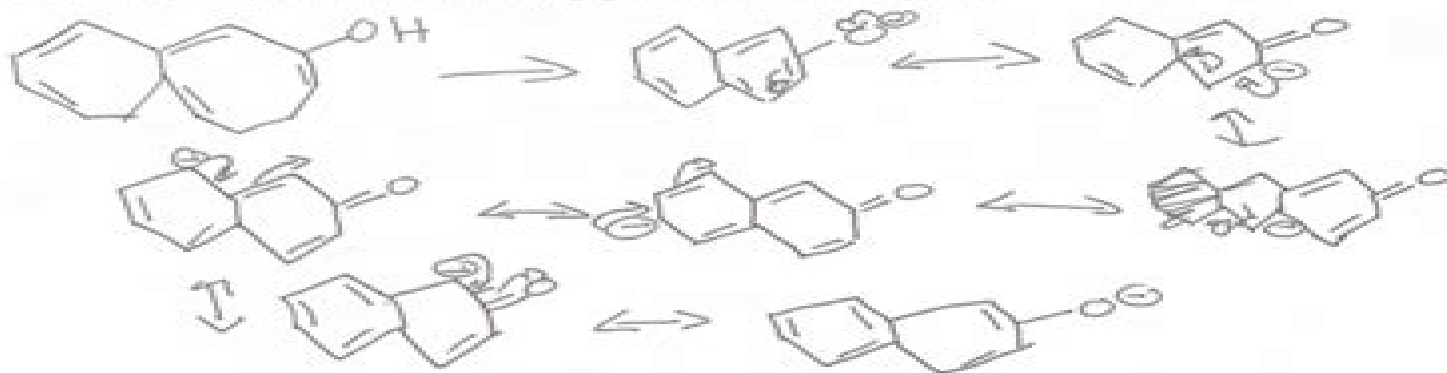
Chemistry 2123 Laboratory: Extractions

Questions:

1. Draw the resonance structures for the benzoate ion.



2. Draw the resonance structures for the conjugate base anion of 2-naphthol.



3. Why do you suppose that the fewer resonance structures for the carboxylate are more effective at increasing the acidity of a carboxylic acid than those of the conjugate base of 2-naphthol?

In benzoate ion, anionic charge is stabilized the most by electronegative oxygen. In the 2-naphthol conjugate base anionic charge is stabilized by the carbon, less stable. Benzoate ion is more stable than 2 naphthol.

4. Why is only benzoic acid extracted into the aqueous phase when using sodium bicarbonate, whereas 2-naphthol required the use of sodium hydroxide?

Benzoic Acid is strong acid + Sodium bicarbonate is a weak base. Moreover, 2-naphthol is a weak acid and sodium hydroxide is a strong base.

Chem 14laboratory Manual Answers

Francis Day Curtis



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