

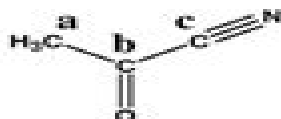
CHEM 241

Homework_Chemical Bonding: Hybridization, Bond Polarity, Molecular Polarity

Using the lecture notes and the provided short videos, please solve the assignments below by typing the answers directly in this document. When you are done, please convert the Word file to pdf and upload it to Turnitin via the provided link by **Tue, June 30, 11:59 PM**.

1.

a) Identify the hybridization of each atom other than H in the structure below:



b) What are the atoms geometry and the bond angle around each of the atoms considered in point a)?

c) What types of orbitals overlap to form the carbon-nitrogen triple bond?

Hint: see the examples solved on class. Please indicate the type of bonds in the triple bond, followed by the corresponding orbital overlap in parenthesis.

Answers:

C^a : tetrahedral, sp^3 , $\sim 109^\circ$

C^b : trigonal planar, sp^2 , $\sim 120^\circ$

O: linear, sp^2 , no bond angle

C^c : linear, sp , $\sim 180^\circ$

N: linear, sp , no bond angle

One sp orbital from C^c overlaps with one sp orbital from N to form the C-N σ ($sp-sp$) bond.

Two p orbitals from C^c overlap with two p orbitals from N to form the two C-N π ($p-p$) bonds.

2. Place the following elements in order of decreasing electronegativity, based on their size. Explain your answers briefly.

S Cl Se

Answer:

Cl > S > Se

Cl is the smallest in volume, hence the valence shell is closer to the nucleus; Cl nucleus can thus attract easier valence electrons from other atoms to form bonds, meaning it would be expected to have the highest electronegativity.

Highlight or underline the correct answer or answers for the questions below:

Chem Fax Chemical Bonding Answers

Annelies Wilder-Smith



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