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Chemistry 12  
**INVESTIGATING CHEMICAL EQUILIBRIUM**

The chemical reactions you have studied to this point have proceeded to completion. However, many reactions proceed only to a state of *equilibrium*, where both reactants and products exist together. This results from some of the product molecules having sufficient energy to reform the reactants in a reverse reaction. A state of equilibrium is established when the rates of the forward and reverse reactions are equal. Although both reactions continue to occur, there is no net change in the observable properties (called *macroscopic properties*) such as concentration, pressure (for gases) or colour intensity of a coloured solution. In a reversible reaction, constant macroscopic properties (including temperature) indicate that equilibrium is present.

Any change to the conditions at equilibrium, such as concentration, pressure or temperature, is said to produce a *stress* on the equilibrium. Le Châtelier's principle allows us to predict the effects of such a stress. The stress produces a shift in the system that attempts to counteract the imposed stress and a new equilibrium is established. The macroscopic properties of the system change, until they become constant once more as the new equilibrium is reached.

You will study three different equilibrium systems involving ions in solution. The first system is the conversion of the indicator bromocresol green from its blue form to its yellow form. Its structure is rather complex so the formula is simplified as HBcg. The extent to which the forward reaction is favoured depends upon the concentration of the hydrogen ions in solution:



The second system involves the light yellow iron(III) ion,  $\text{Fe}^{3+}$ , and colourless thiocyanate ion,  $\text{SCN}^-$ . They react together to form the complex ion  $\text{FeSCN}^{2-}$  which exhibits a blood-red colour:



The effect of changing the concentrations of the reactant ions will be studied.

The last system (Part III) involves the equilibrium between hydrated copper(II) ion,  $\text{Cu}(\text{H}_2\text{O})_6^{2+}$ , and the tetraamminocopper(II) ion,  $\text{Cu}(\text{NH}_3)_4^{2+}$ , in which the water molecules have been replaced with  $\text{NH}_3$  molecules. The effect of acid and of  $\text{NH}_3$  on this equilibrium will be examined.

**OBJECTIVES:**

1. to recognize the macroscopic properties of three chemical systems at equilibrium
2. to observe shifts in equilibrium concentrations as stresses of concentration changes are applied to the systems
3. to explain the observations obtained by applying Le Châtelier's principle

# Chem Fax Lab Answers Chemical Equilibrium

**Katrin Zwirgmaier**



## Chem Fax Lab Answers Chemical Equilibrium:

*Fundamentals of Chemistry: Laboratory Studies* Frank Brescia, 2012-12-02 *Fundamentals of Chemistry Laboratory Studies* Third Edition is a manual that provides instruction on techniques of chemical laboratory operations Each experiment is discussed in terms of the major objective the experimental approach to the objective the measurements or observations to be made and the calculation and interpretation of results Topics covered include manipulation weights and measures molecular weight acids and bases gravimetric and volumetric stoichiometry and thermochemistry This book is comprised of 43 chapters divided into 14 sections and begins by presenting general information on metric and other units common laboratory equipment and chemical laboratory methods The first chapter introduces the reader to the Bunsen burner and the principles of glass working followed by a discussion on mass and volume measurements including the determination of density The following chapters focus on states of matter molecular weight stoichiometry and intermolecular forces Preparations and syntheses are also considered along with chemical equilibrium and electrochemistry The final section is devoted to qualitative analysis particularly of cations and anions This monograph is intended primarily for students of chemistry

*A Laboratory Manual of General Chemistry* Leon Burr Richardson, Andrew Jackson Scarlett, 1929 *The American Catalogue*, 1905 *A Laboratory outline of general chemistry* Alexander Smith, 1899 *Chemical Abstracts*, 1910 **The Journal of Industrial and Engineering Chemistry**, 1914 **Register** Cornell University, 1892 *"The" Cornell University register* Cornell University, 1893 **The American Physician**, 1925 **The Education Index**, 1986 *Chemical News and Journal of Physical Science*, 1895 *Chemical News and Journal of Industrial Science*, 1895 *Books in Print Supplement*, 2002 *Chemical Engineering Catalog*, 1919 **A Dictionary of Applied Chemistry** Sir Thomas Edward Thorpe, 1921 *A Laboratory Outline of College Chemistry* Alexander Smith, 1916 *Current Index to Journals in Education*, 1985 *Laboratory Manual of General Chemistry* Arthur Becket Lamb, 1916 **Journal of Chemical Education**, 1925 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *The Chemical Trade Journal and Chemical Engineer*, 1908

## Adopting the Tune of Phrase: An Psychological Symphony within **Chem Fax Lab Answers Chemical Equilibrium**

In a global taken by screens and the ceaseless chatter of immediate communication, the melodic beauty and emotional symphony produced by the prepared word usually fade in to the back ground, eclipsed by the relentless noise and disruptions that permeate our lives. However, set within the pages of **Chem Fax Lab Answers Chemical Equilibrium** an enchanting literary value filled with fresh thoughts, lies an immersive symphony waiting to be embraced. Crafted by an outstanding composer of language, this interesting masterpiece conducts visitors on a mental journey, well unraveling the concealed tunes and profound impact resonating within each carefully crafted phrase. Within the depths with this touching examination, we shall examine the book is main harmonies, analyze their enthralling writing style, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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