

ChemActivity 2: Resonance Structures

(Why is baking soda a much weaker base than hydroxide, though both have an O with a -1 charge?)

Model 1: Two Lewis Structures for Baking Soda (bicarbonate)

There are two different but equally good ways to draw a legitimate Lewis structure of bicarbonate ion.



Figure 2.1: Bicarbonate Ion (HCO_3^-)

To facilitate the following discussion, the three oxygens of bicarbonate were labeled O_1 , O_2 , and O_3 .

Neither structure by itself tells the whole story.

To get the most accurate picture of bicarbonate ion you must average the two different structures.

Construct Your Understanding Questions (to do in class)

1. A student looks at Model 1 and says she expects the bond between C and O_1 to be halfway between a single bond and a double bond (i.e., shorter and stronger than a single bond, but not as short and strong as a double bond) and that the same will be true of the bond between C and O_2 .

Use the information in Model 1 to construct an explanation for these conclusions.

2. Based on the most accurate picture of bicarbonate ion, what is the charge on O_1 ? ... O_2 ? ... O_3 ? (Note that it may be a fraction.) Briefly explain how you calculated your answers.
3. Organic chemists say that "the -1 charge on bicarbonate ion is 'delocalized' between O_1 and O_2 ." Come up with at least one synonym for the word "delocalized" in the preceding sentence.
4. Which do you expect to have a more intense and concentrated "hot spot" of negative charge: methoxide ion or bicarbonate ion? Explain.
5. One definition of a strong base is: "A molecule with a -1 formal charge localized on a single H, N, or O atom." Construct an explanation for why bicarbonate ion is NOT a strong base.



Chemactivity 2 Answer Keys

Seth Crawley



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Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Prentice Hall Chemistry*, 2000

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