

Ions and Ionic Compounds

(How are ionic compounds formed?)

Model 1: Common charges (by group) on elements when in ionic compounds.

Charge		← variable charge →																			
Grp #																					
1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18				
1																			2		
H																			He		
3	4											5	6	7	8	9	10				
Li	Be											B	C	N	O	F	Ne				
11	12											13	14	15	16	17	18				
Na	Mg											Al	Si	P	S	Cl	Ar				
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36				
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr				
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54				
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe				
55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86				
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn				
87	88	89	104	105	106	107	108	109	110	111	112		114		116						
Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn		Fl		Lv						

A cation has a positive charge.

An anion has a negative charge.

Critical Thinking Questions:

- Recall the shell model of the atom.
 - How many electrons fill the first shell? _____
 - How many electrons fill the second shell? _____
- How many valence electrons are found in atoms in
 - group 1? _____
 - group 13? _____
 - group 16? _____
- When atoms form ionic compounds, they often fill their outer valence shell (**octet rule**) by either gaining or losing 1, 2, or 3 electrons. Based on your answer to CTQ 2, indicate how many electrons these atoms would either gain or lose to achieve an octet.
 - group 1? _____
 - group 13? _____
 - group 16? _____
- What would be the resulting charge on the ions created in CTQ 3 when electrons are gained or lost?
 - group 1? _____
 - group 13? _____
 - group 16? _____
- Based on your answer to CTQ 4, do metals typically form **anions** or **cations** (circle one)?
- Work with your team and write charges in the empty boxes above each group (skip the transition metals, groups 3-12) at the top of the periodic table in Model 1.
- Do nonmetals typically form **anions** or **cations** (circle one)?

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**Michael P. Garoutte, Ashley B.
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