

ChemActivity 8

Acid-Base Reactions

(How can you predict if an acid-base reaction will be up-hill or down-hill?)

Model 1: Brønsted-Lowry Acids and Bases

Figure 1a: Examples of Acid/Base Conjugate Pairs

- **ACID:** has an H^+ to give (proton donor) (H^+ = a proton)
- **BASE:** can make a bond to H^+ (proton acceptor)
- The stronger the conjugate acid, the weaker its conjugate base.

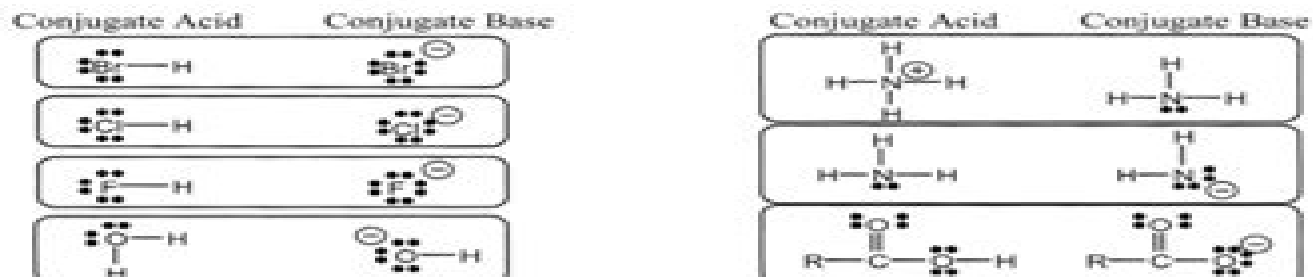


Figure 1b: Brønsted-Lowry Acid-Base Reaction ("Proton Exchange")



Critical Thinking Questions

1. Write the reverse of the reaction in Figure 1b. Show the forming and breaking of bonds using curved arrows.

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L Towne

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the 'L Towne' text.

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Enjoying the Beat of Appearance: An Psychological Symphony within **Chemactivity 8 Answers**

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