

## CHEM 108

### ChemActivity 15: Balancing Chemical Equations and Percent Composition Solution Key

#### The Model: Writing Chemical Equations to be Balanced

*Fact:* Trihydrogen mononitride reacts with oxygen to yield nitrogen monoxide and dihydrogen monoxide.

With our understanding of nomenclature, we may begin to write the following (as-yet unbalanced) chemical equation to describe the reaction.



The **reactants** are separated from the **products** by an arrow ( $\rightarrow$ ), which points from the reactants to the products. The reactants are separated from each other by + signs, as are the products. Because it violates the **Law of Conservation of Matter** (3 H atoms  $\neq$  2 H atoms), the above is not yet balanced. To balance it, we have yet to write the (non-1) **stoichiometric coefficients** (numbers placed in front of each reactant's/product's formula).

#### Key Questions

1. Consider the chemical equation and the corresponding microscale representations of the reaction involving Lewis structures. Both " $4\text{P}$ " and " $\text{P}_4$ " involve four phosphorus atoms, but how do the notations " $4\text{P}$ " and " $\text{P}_4$ " differ?



" $4\text{P}$ " implies there are four unbound phosphorus atoms, while " $\text{P}_4$ " implies there is a molecule consisting of four phosphorus atoms.

2. To the right is a balanced chemical equation.



But this is *not* the correct way to balance the equation shown in the Model. (Why not?)

The formula of water is  $\text{H}_2\text{O}$ , not  $\text{H}_2\text{O}_2$ . One cannot change the formula of a reactant or product in order to balance a chemical equation.

When we balance chemical equations, we introduce new subscripts to chemical formulas / coefficients that precede chemical formulas. (Circle your answer.)

# Chemactivity 10 Key

**CL Gary**



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