CHEM 108 ChemActivity 16: Stoichiometry of Chemical Equations Solution Key

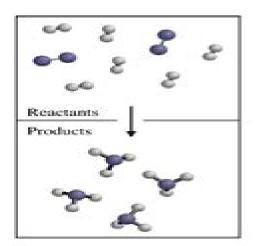
The Model

The reaction between elemental hydrogen and elemental nitrogen to produce ammonia is given by eq 1.

$$3 H_2(g) + N_2(g) \xrightarrow{Fe} 2 NH_3(g)$$
 (1)

The chemical equation can be interpreted in two different ways:

- Microscopically Three molecules of hydrogen react with one molecule of nitrogen to yield two molecules of ammonia.
- (2) Macroscopically For every three moles of hydrogen that react, one mole of nitrogen reacts to produce two moles of ammonia.



Key Questions

1. While describing the reaction on a "microscopic scale", why would it be inappropriate to use the following?

$$^{3}/_{2} \text{ H}_{2}(g) + ^{1}/_{2} \text{ N}_{2}(g) \rightarrow \text{ NH}_{3}(g)$$

On a microscopic scale, it would be "three-halves of a hydrogen molecule react with half a nitrogen molecule to yield one ammonia molecule." But there is no way to have half of a molecule! (Monatomic N doesn't react the same as N₂.)

2a. In the synthesis of ammonia from its elements, how many moles of H2 are consumed for every mole of N2 that is consumed?

In eq 1, the stoichiometric coefficients in front of N₂ and H₂ are 1 and 3, respectively. So 3 moles of H₂ will be consumed for every 1 mole of N₂ that is consumed.

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