

Acid-Base Titrations Prelab

1. Determine the volume released in the burets on the right. Remember significant figures and units!

Initial: 0.90 mL

Final: 20.21 mL

19.31 mL

2. If the concentration of the solution (NaOH) is 0.1500 M, calculate the number of moles of NaOH in the volume in (1). Show your work.

Moles of solute = concentration X volume

Volume used: final reading – initial reading

= 20.21 mL – 0.90 mL = 19.31 mL

Volume used = 19.31 mL / 1000 mL/L = 0.01931 L

Moles of NaOH: concentration X volume used

= 0.1500 M X 0.01931 L

= 0.0028965

Initial Reading



Final Reading



(For the following questions, please refer to the report on Blackboard.)

1. What is the concentration of the NaOH solution in the buret?  
Answer: 0.1500 M
2. What is the volume of the NaOH solution in the buret?  
Answer: 19.31 mL
3. What is the volume of the NaOH solution in the buret?  
Answer: 19.31 mL
4. What is the concentration of the NaOH solution in the buret?  
The concentration of the NaOH solution in the buret is 0.1500 M. This is the concentration of the NaOH solution in the buret. The concentration of the NaOH solution in the buret is 0.1500 M. The concentration of the NaOH solution in the buret is 0.1500 M.

# Chemfax Acid Base Titrations Lab Prelab Answers

**Sten Johansson**



## **Chemfax Acid Base Titrations Lab Prelab Answers:**

Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz, 1973      Identification of Misconceptions Concerning Acid-base Titrations Lockwood Ryerson Gibbs, 1995      *Aqueous Acid-base Equilibria and Titrations* Robert De Levie, 1999

This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry. It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way. This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems; these approaches are often simpler than the approximations which they replace. Acid-base problems are readily addressed in terms of the proton condition, a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical approach of Bjerrum, Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria. Based on these concentrations, the proton condition can usually be simplified, often leading directly to the value of the pH. In the description of acid-base titrations, a general master equation is developed. It provides a continuous and complete description of the entire titration curve, which can then be used for computer-based comparison with experimental data. Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated. Activity effects are described in detail, including their effect on titration curves. The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements, which are subject to activity corrections, and balance equations and spectroscopic pH measurements, which are not. Finally, an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved.

**Acid/Base Titrations** Michael O'Brien, 1999      *General Expressions for Acid-Base Titrations of Arbitrary Mixtures*, 1995. A single general master equation is given for acid-base titrations, describing the entire progress of the titration and equally valid for the titration of a strong acid with a strong base as for that of the titration of an arbitrary mixture of acids with an arbitrary mixture of bases or vice versa.

**pH, Buffers and Acid-base Titrations** LaMotte Chemical Products Co. Educational Products Division, 1977?      *Numerical Evaluation of Potentiometric Acid-base Titrations* Sten Johansson, 1978

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