

Pre-Lab Questions (Use a separate sheet of paper to answer the following questions.)

1. In Part 1 of this laboratory, 25.0 mL of a 0.100 M solution of the weak acid acetic acid, CH_3COOH , is titrated with a 0.100 M solution of the strong base sodium hydroxide, NaOH . K_a of acetic acid is 1.8×10^{-5} .

a. Calculate the pH of the equivalence point. Enter this value in the Part 1 Data Table.

b. Five indicators, along with their pH , pK_a and their color change are shown below.

$5.6 \times 10^{-10} \approx \frac{x^2}{0.050}$ $x = 5.3 \times 10^{-6}$ $\text{pOH} = 5.28$ $\text{pH} = 8.72$

	pK_a	1	2	3	4	5	6	7	8	9	10	11	12	13
Methyl Orange	3.40	Red												
Methyl Red	4.95	Red												
Phenolphthalein	9.4				Colorless								Red	
Bromothymol Blue	7.1			Yellow								Blue		
Thymolphthalein	10.0													

$\text{CH}_3\text{COO}^- + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COOH} + \text{OH}^-$

I Based on the pH calculated in 1a., select the appropriate indicator for the weak acid-strong base titration. Enter the selection in the Part 1 Data Table.

C 2. In Part 2, 25.0 mL of a 0.10 M solution of the weak base ammonia, NH_3 , is titrated with a 0.10 M solution of the strong acid hydrochloric acid, HCl . K_b of ammonia is 1.8×10^{-5} .

E a. Calculate the pH of the equivalence point. Enter this value in the Part 2 Data Table.

b. Select an indicator for the weak base-strong acid titration. Enter the selection in the Part 2 Data Table.

Chemfax Lab Report Answers

Valerie Ludwig Lechtanski



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