

AP CHEMISTRY LAB LIMITING REACTANT

Objectives:

- To determine the limiting reactant in a salt mixture.
- To determine the percent composition of each substance in a salt mixture.

Introduction

Two factors limit the yield of products in a chemical reaction: (1) the amounts of starting materials (reactants) and (2) the percent yield of the reaction. Many experimental conditions, for example, temperature and pressure, can be adjusted to increase the yield, especially of a desired product, but because chemicals react according to fixed mole ratios (stoichiometrically), only a limited amount of product forms from given amounts of starting materials. The reactant determining the amount of product generated in a chemical reaction is called the **limiting reactant** in the chemical system.

To better understand the concept of the limiting reactant, let's look at the reaction that is under investigation in this experiment. The reaction of sodium phosphate dodecahydrate, $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$ and barium chloride dihydrate, $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$. The molecular equation is:



In this experiment an unknown salt mixture of $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$ and $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ is added to water producing insoluble $\text{Ba}_3(\text{PO}_4)_2$. The mass of the initial salt mixture and the mass of the final product, $\text{Ba}_3(\text{PO}_4)_2$, are measured. A series of tests is also used to determine the limiting reactant. Finally, the percent composition of the original salt mixture is calculated.

Pre-laboratory Questions:

1. The solubility of $\text{Ba}_3(\text{PO}_4)_2$ is 0.519 mg/L. How many milligrams and moles of $\text{Ba}_3(\text{PO}_4)_2$ dissolve in 200. ml of solution?
2. Describe how the mass of $\text{Ba}_3(\text{PO}_4)_2$ precipitate is determined in Part A.
3. Describe how to test your unknown salt mixture for the presence of excess $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}$.
4. Why is fine porosity paper required for use in the filtering of the $\text{Ba}_3(\text{PO}_4)_2$ precipitate?

Materials: Salt mixture, DI water, small beakers, scoop, electronic balance, rinse bottle, filter paper, funnel, stir rod, watch glass, tongs, hot plate, 1000 mL beaker, Erlenmeyer flask, dropper bottle 0.50 M $\text{Na}_3\text{PO}_4 \cdot 12\text{H}_2\text{O}_{\text{aq}}$, dropper bottle of 0.50 M $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}_{\text{aq}}$.

Procedure:

Two trials are recommended for this experiment. To hasten the analysis, weigh duplicate unknown salt mixtures and simultaneously follow the procedure for each. Label the beakers accordingly for trial 1 and trial 2 to avoid the intermixing of samples and solutions.

A. Precipitation of $\text{Ba}_3(\text{PO}_4)_2$ from the salt mixture

1. Weigh and Dissolve the Salt Mixture. Weigh (± 0.001 g) on weighing paper about 1.0 gram of the unknown salt mixture. Record the mass for trial 1 in your laboratory notebook. Repeat for trial 2. Transfer the salt mixture to a labeled 400-ml beaker and add 200-ml of deionized water. Stir the mixture with a stirring rod for about 1 minute and then allow the precipitate to settle. Leave the stirring rod in the beaker.
2. Digest the Precipitate. Cover the beaker with a watchglass and maintain the solution hot ($80\text{--}90^\circ\text{C}$) on a steam bath or with a **low** flame for 30 minutes. While the precipitate is being warmed, proceed to part A.3. Periodically check on the progress of heating. After 30 minutes, allow the precipitate to settle; it does not need to be cooled to room temperature.
3. Filter the Precipitate. Set up a gravity filtering apparatus. Weigh (± 0.001 g) a piece of Whatman No.42 or Fisherbrand Q2 filter paper fine porosity filter paper. Fold it properly and seal it into the filter funnel with a small amount of deionized water. Discard the water in the receiving flask. Return to Part A.2.
4. Separation of Supernatant in Part A.2. Once the supernatant has cleared in Part A.2, decant two 50-ml volumes into separate 100-ml beakers, labeled Beaker I and Beaker II. Save for part B.
5. Filter the Precipitate Salt. While the solution is still warm from Part A.4, quantitatively transfer the precipitate to the filter funnel. Remove any precipitate on the wall of the beaker with a rubber policeman, and hot ($80\text{--}90^\circ\text{C}$) water*, wash the precipitate onto the filter. Rinse the $\text{Ba}_3(\text{PO}_4)_2$ precipitate with two 5-ml portions of hot water (see Part A.7).
*I'll keep a hot water bath up front for this purpose.
6. Dry and Weigh the Salt Precipitate. Remove the filter paper holding the precipitate from the filter funnel. Air-dry the precipitate on the filter paper until the next laboratory period or in a 110°C constant-temperature drying oven overnight. Weigh (± 0.001 g) the precipitate and filter paper. Record in notebook.

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