

Name: _____



Balancing Chemical Equations

Directions: Balance each of the following chemical equations.

- 1) ____ $C_{30}H_{42} + O_2 \rightarrow CO_2 + H_2O$
- 2) ____ $Al(OH)_3 + HBr \rightarrow AlBr_3 + H_2O$
- 3) ____ $CH_3CH_2CH_2CH_3 + O_2 \rightarrow CO_2 + H_2O$
- 4) ____ $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
- 5) ____ $Li + AlCl_3 \rightarrow LiCl + Al$
- 6) ____ $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$
- 7) ____ $NH_4OH + H_3PO_4 \rightarrow (NH_4)_3PO_4 + H_2O$
- 8) ____ $Rb + P \rightarrow Rb_3P$
- 9) ____ $CH_4 + O_2 \rightarrow CO_2 + H_2O$
- 10) ____ $Al(OH)_3 + H_2SO_4 \rightarrow Al_2(SO_4)_3 + H_2O$
- 11) ____ $Na + Cl_2 \rightarrow NaCl$
- 12) ____ $Rb + S_8 \rightarrow Rb_2S$
- 13) ____ $H_3PO_4 + Ca(OH)_2 \rightarrow Ca_3(PO_4)_2 + H_2O$
- 14) ____ $NH_3 + HCl \rightarrow NH_4Cl$
- 15) ____ $Li + H_2O \rightarrow LiOH + H_2$
- 16) ____ $Ca_3(PO_4)_2 + SiO_2 + C \rightarrow CaSiO_3 + CO + P$
- 17) ____ $NH_3 + O_2 \rightarrow N_2 + H_2O$
- 18) ____ $FeS_2 + O_2 \rightarrow Fe_2O_3 + SO_2$

Chemfiesta Chemical And Physical

Saul I. Siegal



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Physical Properties of Chemical Compounds, 1955

PHYSICAL PROPERTIES OF CHEMICAL COMPOUNDS

Robert R. Dreisbach, 1980 *Advances in chemistry : physical properties of chemical compounds - III* Robert Rickert

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Constitution and Some Physical Properties Samuel Smiles, 1910 **Chemical Reactivity in Liquids** Michael

Moreau, Pierre Turq, 2012-12-06 Understanding chemical reactivity has been the permanent concern of chemists from time immemorial If we were able to understand it and express it quantitatively there would practically remain no unsolved mystery and reactions would be fully predictable with their products and rates and even side reactions The beautiful developments of thermodynamics through the 19th century supplied us with the knowledge of the way a reactions progresses and the statistical view initiated by Gibbs has progressively led to an unders tanding closer to the microscopic phenomena But is was always evident to all that these advances still left our understanding of chemical reactivity far behind our empirical knowledge of the chemical reaction in its practically infinite variety The advances of recent years in quantum chemistry and statistical mechanics enhanced by the present availability of powerful and fast compu ters are very fast changing this picture and bringing us really close to a microscopic understanding of chemical equilibria reaction rates etc This is the reason why our Society encouraged a few years ago the initiative of Professor Savo Bratos who with a group of French colleagues prepared an impressive study on Reactivite chimique en phase liquide a prospective report which was jointly published by the Societe Fran **The Relations between chemical constitution and some physical properties**

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The Relations

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