

100

27. It is necessary to know the formula of a compound in order to calculate its percentage composition.

AT

28. If the percent by mass of carbon in methane,  $\text{CH}_4$ , is 75%, then 100 grams of methane contains 75 g of percent of hydrogen.

AT

29. The formula for methane,  $\text{CH}_4$ , is both a molecular and an empirical formula.

100

30. The empirical formula of glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , is  $\text{CH}_2\text{O}$ .

C. Identification

Write the following particles in the space provided. Show your work.

31. How many neutrons are in 5.1 g of  $^{235}\text{U}$  versus 5.1 g?

32. How many representative particles are in 2.18 g  $^{235}\text{U}$  and  $\text{H}_2\text{O}$ ?

33. What is the gram molecular mass of methane,  $\text{CH}_4$ ?

34. Find the mass of each of the following:

a. 3.45 g  $^{235}\text{U}$  and  $\text{H}_2\text{O}$

b. 2.43 g and  $\text{H}_2\text{O}$

c. 4.50 g and 1.30 g

d. 0.78 g and 0.2

# Chemfiesta Chemical Quantities Answer Key

**BM King**



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