

LEARN Science • Chemistry  
Balancing Chemical Equations

Part 1: Balance these chemical equations. Record what type of reaction you think is the likely result in the right margin.

1. $\text{BaCl}_2 + \text{AgNO}_3 \rightarrow \text{Ba(NO}_3)_2 + \text{AgCl}$	DC
2. $\text{Fe(OH)}_3 \rightarrow \text{Fe}_2\text{O}_3 + \text{H}_2\text{O}$	D
3. $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$	SR
4. $\text{Pb(NO}_3)_2 + \text{K}_2\text{CrO}_4 \rightarrow \text{PbCrO}_4 + 2\text{KNO}_3$	DP
5. $\text{Al}_2(\text{SO}_4)_3 + \text{Zn(OH)}_2 \rightarrow 2\text{Al(OH)}_3 + \text{Zn}_2(\text{SO}_4)_3$	DP
6. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$	SR
7. $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$	SR
8. $2\text{Fe}_2\text{O}_3 \rightarrow 4\text{Fe} + 3\text{O}_2$	D
9. $2\text{HBr} + \text{Ba}_2\text{O} \rightarrow 2\text{BaBr}_2 + \text{H}_2\text{O}$	SR
10. $\text{Cr}_2\text{O}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CrO}_4$	S
11. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	S
12. $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH}$	S

Part 2: Write balanced chemical equations that describe the reactions below. Remember that all acids completely ionize in water (aqueous), and/or are highly polarized when appropriate.

1. Two metal-oxides react to produce a single metal oxide solid.



2. Chromium(III) and sulfate react with  $\text{Ba}^{2+}$  to form barium sulfate (the pattern sulfate and hydrogen peroxide).



3. Barium chloride reacts with aqueous hydrochloric acid (HCl) to give product barium chloride solid and hydrogen chloride gas.



# Chemfiesta Nuclear Chemistry Answers

**Ian Guch**



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**Handbook of Nuclear Chemistry** Attila Vértés, Sándor Nagy, Zoltán Klencsár, Rezso György Lovas, Frank Rösch, 2010-12-10 This revised and extended 6 volume handbook set is the most comprehensive and voluminous reference work of its kind in the field of nuclear chemistry The Handbook set covers all of the chemical aspects of nuclear science starting from the physical basics and including such diverse areas as the chemistry of transactinides and exotic atoms as well as radioactive waste management and radiopharmaceutical chemistry relevant to nuclear medicine The nuclear methods of the investigation of chemical structure also receive ample space and attention The international team of authors consists of scores of world renowned experts nuclear chemists radiopharmaceutical chemists

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