

## pH Calculations – Answers

- 1) Determine the pH of the following solutions:
- a) A  $4.5 \times 10^{-3}$  M HBr solution.  
**2.4**
  - b) A  $3.67 \times 10^{-5}$  M KOH solution.  
**9.56**
  - c) A solution made by diluting 25 mL of 6.0 M HCl until the final volume of the solution is 1.75 L.  
**1.1 (the diluted solution is 0.086 M HCl)**
  - d) 5 L of an aqueous solution that contains 1.0 grams of HBr and 1.0 grams of nitric acid.  
**2.2 (the solution has an overall acid concentration of 0.0056 M)**
- 2) What are the pOHs for the solutions in problem 1?
- a) **11.6**
  - b) **4.44**
  - c) **12.9**
  - d) **11.8**
- 3) Explain why even a basic solution contains some  $\text{H}^+$  ions.  
**All aqueous solutions contain  $\text{H}^+$  ions from the autoionization of water,  $\text{H}_2\text{O} \rightarrow \text{H}^+ + \text{OH}^-$ .**
- 4) Explain why even an acidic solution contains some  $\text{OH}^-$  ions.  
**The same answer from #3 applies here.**
- 5) More challenging: What is the pH of a  $1.5 \times 10^{-10}$  M HBr solution?  
**If you do the calculation to find  $-\log(1.5 \times 10^{-10})$ , you get an answer of 9.8. However, this is intuitively incorrect – after all, how can a solution that's made of nothing but pure water (pH = 7) with an acid added be basic overall? The answer, it can't. The actual pH of the solution is just about 7, with the main acid source being the  $\text{H}^+$  formed from the autoionization of water.**
- 6) Does your answer from problem #5 make sense? Explain.  
**See answer from #5.**

# Chemfiesta Ph Answers

**Ian Guch, Kjirsten Wayman Ph.D.**



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