



# Atom, molecules and Chemical Arithmetic

Chemistry is basically an experimental science. In it we study physical and chemical properties of substance and measure it upto possibility. The results of measurement can be reported in two steps : (a) Arithmetic number, (b) Unit of measurement.

Every experimental measurement varies slightly from one another and involves some error or uncertainty depending upon the skill of person making the measurements and measuring instrument. The closeness of the set of values obtained from identical measurement is called **precision** and a related term, refers to the closeness of a single measurement to its true value called **accuracy**.

## 1.1 Significant figures

In the measured value of a physical quantity, the digits about the correctness of which we are surplus the last digit which is doubtful, are called the significant figures. Number of significant figures in a physical quantity depends upon the least count of the instrument used for its measurement.

(1) **Common rules for counting significant figures** : Following are some of the common rules for counting significant figures in a given expression:

**Rule 1.** All non zero digits are significant.

Example :  $x = 1234$  has four significant figures. Again  $x = 189$  has only three significant figures.

**Rule 2.** All zeros occurring between two non zero digits are significant.

Example :  $x = 1007$  has four significant figures. Again  $x = 1.0809$  has five significant figures.

**Rule 3.** In a number less than one, all zeros to the right of decimal point and to the left of a non zero digit are not significant.

Example :  $x = 0.0084$  has only two significant digits. Again,  $x = 1.0084$  has five significant figures. This is on account of rule 2.

**Rule 4.** All zeros on the right of the last non zero digit in the decimal part are significant.

Example :  $x = 0.00800$  has three significant figures 8, 0, 0. The zeros before 8 are not significant again 1.00 has three significant figures.

**Rule 5.** All zeros on the right of the non zero digit are not significant.

Example :  $x = 1000$  has only one significant figure. Again  $x = 37800$  has three significant figures.

**Rule 6.** All zeros on the right of the last non zero digit become significant, when they come from a measurement.

Example : Suppose distance between two stations is measured to be 3050 m. It has four significant figures. The same distance can be expressed as 3.050 km or  $3.050 \times 10^5$  cm. In all these expressions, number of significant figures continues to be four. Thus we conclude that

# Chemical Arithmetic Atoms And Molecules 1

**James E. Brady, Gerard E. Humiston**



## **Chemical Arithmetic Atoms And Molecules 1:**

**The Elements of Chemical Arithmetic** J. Milnor Coit, 1886      **Chemistry For Aieeee** , *Comprehensive Objective Book For Aieeee* Narinder Kumar, Dr. J. K. Juneja, Dr. N. K. Verma, Dr. Kulbhushan Prakash, 2006      Iit-Jee Main and Advanced Chemistry Dr. K. G. Ojha, 2022-07-02 The new edition of IIT JEE Main Advanced CHEMISTRY is designed to present a whole package of Chemistry study preparation sufficing the requirements of the aspirants who are preparing for the upcoming exam Highlights of the Book Exam Pattern and Chemistry Syllabus for JEE Main and Advanced included An Analysis of IIT JEE included Chapter wise Theory detailed with 1000 examples 5000 Chapter wise Multiple Choice Questions 2500 Chapter wise Different Format Questions Chapter wise Assessment Test Chapter wise HOTS Problems Appendix on Equations Glossary JEE Main and Advanced Mock Test NEET Mock Test Answers to Questions included with Explanations Presence of accurate Diagrams and Tables From food to pharmaceuticals Chemistry plays a huge role in making informed decisions Therefore this book proves a comprehensive resource of Chemistry and serves to be a suitable Study Guide for the aspirants with focus on Qualitative Preparation and Systematic understanding of the Syllabus and Examination Level With provision for self assessment in Mock Tests this book stands beneficial in imprinting concepts in the mind      Science and Art of Mining , 1921      *Chemistry in Quantitative Language* Christopher O. Oriakhi, 2009 Problem solving is one of the most challenging aspects students encounter in general chemistry courses leading to frustration and failure Consequently many students become less motivated to take additional chemistry courses after their first year This book deals with calculations in general chemistry and its primary goal is to prevent frustration by providing students with innovative intuitive and systematic strategies to problem solving in chemistry The material addresses this issue by providing several sample problems with carefully explained step by step solutions for each concept Key concepts basic theories and equations are provided and worked examples are selected to reflect possible ways problems could be presented to students      General Chemistry James E. Brady, Gerard E. Humiston, 1975      **Chemical primer** Simeon Pease Meads, 1884      **Chemistry, metallurgy, engines** American School (Lansing, Ill.), 1903      *Qualitative Chemical Analysis* Albert Benjamin Prescott, Otis Coe Johnson, 1890

**Cyclopedia of Heating, Plumbing and Sanitation: Hydraulics. Water supply. Chemistry. Bacteriology and sanitation** , 1909      *Elements of Chemistry* Simeon P. Meads, 1891      Essentials of Modern Chemistry Charles Elwood Dull, 1918      **Canadian Chemical Processing** , 1948      *Fundamentals of Physics, Chemistry and Bacteriology in Agriculture* Frank William Miller, 1914      **Beginning Chemistry** Gustav Ludwig Fletcher, Herbert O. Smith, Benjamin Harrow, 1929      Principles of pharmacy Henry Vinecome Arny, 1917      *Proceedings of the High School Conference of ...* , 1928      *Proceedings of the High School Conference of November 1910-November 1931* , 1929      **The Physical Sciences** Frank W. Dobbs, Albert William Forslev, Robert L. Gilbert, 1972

## Decoding **Chemical Arithmetic Atoms And Molecules 1**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chemical Arithmetic Atoms And Molecules 1**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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