

CHEMICAL EQUILIBRIUM MULTIPLE CHOICE QUESTIONS

Read the questions properly before answering.

1. It is a condition that occurs when opposing reactions are proceeding at equal rates.
 - a. Chemical kinetics
 - b. Chemical equilibrium
 - c. Chemical reaction
 - d. Chemical equation
2. Which of the following is not included when writing equilibrium constant expression?
 - a. gases
 - b. aqueous solutions
 - c. liquids
 - d. pure liquids and solids
3. Equilibrium constant greater than 1 means
 - a. equilibrium constant lies to the right , products predominate.
 - b. equilibrium constant lies to the left, reactants predominate.
 - c. equilibrium constant lies to the right, reactants predominate.
 - d. equilibrium constant lies to the left, products predominate.
4. What does it mean when the $K_{eq} < 1$?
 - a. Equilibrium constant lies to the right , products predominate.
 - b. Equilibrium constant lies to the left, reactants predominate.
 - c. Equilibrium constant lies to the right, reactants predominate.
 - d. Equilibrium constant lies to the left, products predominate.
5. What is the equilibrium constant expression of $aA + bB \rightleftharpoons cC + dD$?

a. $K_{eq} = \frac{[a]^a [b]^b}{[c]^c [d]^d}$	c. $K_{eq} = \frac{[C]^c [D]^d}{[a]^a [b]^b}$
b. $K_{eq} = \frac{[A]^a [B]^b}{[C]^c [D]^d}$	d. $K_{eq} = \frac{[C]^c [D]^d}{[A]^a [B]^b}$
6. Which of the following is not a factor that affects chemical equilibrium?
 - a. concentration
 - b. temperature
 - c. pressure and volume
 - d. density
7. Equilibria in which all substances are in the same phases are called
 - a. homogeneous equilibria.
 - b. heterogeneous equilibria.
 - c. equilibrium constant
 - d. equilibrium constant expression.
8. Equilibria in which two or more phases are present in the reaction are called
 - a. homogeneous equilibria.
 - b. heterogeneous equilibria.
 - c. equilibrium constant
 - d. equilibrium constant expression.

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