

Section 1.1

- 1 **a** 2.0 **b** 5.3 **c** 1.3 **d** 10.0
e 5.0 **f** 40.2
2 **a** 144 **b** Neodymium
3

Mass of sample/g	Amount of sample/mol	Number of atoms
—	—	6.02×10^{23}
—	2.00	—
55.8	—	6.02×10^{23}
80.2	—	12.04×10^{23}
63.5	0.50	—

- 4 **a** The mass of the sample is needed to be sure that iodine and oxygen are the only elements in the compound.
b The relative number of moles of iodine and oxygen.
c To change the relative number of moles into the ratio of moles of oxygen relative to 1 mole of iodine.
d In order to produce a ratio involving whole numbers.
e I_2O_5 , I_2O_{10} , I_2O_{20} etc.
f The molar mass is needed.
5 **a** 1 **b** 0.5 **c** 0.25 **d** 0.1
e 0.25 **f** 0.5 **g** 0.25 **h** 0.1
i 2 **j** 5

Atoms of copper are approximately twice as heavy as atoms of sulfur. Thus the same mass contains only half as many moles of copper as it does of sulfur.

- 6 Black copper(II) oxide (CuO) contains equal numbers of copper and oxygen particles (Cu^{2+} and O^{2-} ions). Red copper(I) oxide (Cu_2O) contains twice as many copper particles as oxygen particles (Cu^+ and O^{2-} ions).
7 **a** H_2O **b** CO **c** CS_2 **d** CH_4
e Fe_2O_3 **f** CuO **g** CaO **h** SO_2
i MgH_2
8 **a** 92.5 **b** 7.7 **c** CH_4 **d** MgO
9 **a** SnH_4 **b** CO **c** CO_2 **d** NaHCO_3
e $\text{C}_2\text{H}_6\text{O}$ **f** CaCO_3 **g** HClO_4 **h** NaHCO_3
10 **a** CH_4 **b** P_2O_5 **c** AlCl_3 **d** BH_3
e C_2H_6 **f** C_2H_4 **g** CH_3O **h** $\text{C}_{10}\text{H}_{16}\text{O}_{11}$
11 **a** H_2O_2 **b** CO **c** C_2H_6 **d** C_2H_4
e C_2H_{10}
12 **a** 2 **b** 11 **c** 2 **d** 10
e 2
13 **a** 50.0 **b** 78.0 **c** 129.9 **d** 100.1
e 158.0 **f** 241.8 **g** 132.0
14 **a** 2 **b** 4 **c** 10 **d** 0.02
e 5 **f** 1×10^6

Section 1.2

- 1 **a** $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
b $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
c $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$
d $\text{CaO} + 2\text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$
e $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$
f $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$
g $2\text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$
h $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
i $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
j $2\text{CH}_3\text{OH} + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 4\text{H}_2\text{O}$
2 **a** $2\text{Ca} + \text{O}_2 \rightarrow 2\text{CaO}$
b $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2 + \text{H}_2$
c $\text{C} + \text{CO}_2 \rightarrow 2\text{CO}$
d $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
e $\text{C}_2\text{H}_6 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$
3 **a** $\text{Zn(s)} + \text{H}_2\text{SO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{H}_2(\text{g})$
b $\text{Mg(s)} + 2\text{HCl(aq)} \rightarrow \text{MgCl}_2(\text{aq}) + \text{H}_2(\text{g})$
c $\text{MgCO}_3(\text{s}) \rightarrow \text{MgO(s)} + \text{CO}_2(\text{g})$
d $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O(l)}$
e $\text{BaO(s)} + 2\text{HCl(aq)} \rightarrow \text{BaCl}_2(\text{aq}) + \text{H}_2\text{O(l)}$

Section 1.3

- 1 **a** All the magnesium reacts.
b So that we know the number of moles of each substance involved in the reaction.
c Mass of 1 mole of magnesium oxide.
d Because 2 moles of magnesium oxide are produced.
e To find the mass of magnesium oxide produced from 1g of magnesium.
f $80.6/48.6$ would be multiplied by 50 rather than by 6.
2 20g
3 **a** 2.8g **b** 5.1g **c** 2.5g
4 3667g (3.667 kg)
5 **a** $\text{C}_{10}\text{H}_{16} + 12.5\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$
b 175kg
c 154kg
6 **a** 56 tonnes
b $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
c 64g
d 64 tonnes
e 2 tonnes
f 112 tonnes
7 **a** 217 tonnes
b 0.58 tonnes
8 **a** $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
b 159.6g
c 1.45g
d 1.45 tonnes
e 2.86 tonnes
f 11.9 tonnes

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ML Morrison



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