



# Study Packet: Chemical Names & Formulas

## Set I: Naming or writing formulas of Molecular Compounds

| <i>Name</i>                   | <i>Formula</i>          |
|-------------------------------|-------------------------|
| 1. Dinitrogen Pentoxide       | $\text{N}_2\text{O}_5$  |
| 2. hydrochloric acid          | $\text{HCl}$            |
| 3. Calcium iodide             | <i>ionic</i>            |
| 4. Iodine trichloride         | $\text{ICl}_3$          |
| 5. Sulfur dioxide             | $\text{SO}_2$           |
| 6. Carbon Disulfide           | $\text{CS}_2$           |
| 7. sulfuric acid              | $\text{H}_2\text{SO}_4$ |
| 8. Aluminum oxide             | <i>ionic</i>            |
| 9. Nitrogen Trifluoride       | $\text{NF}_3$           |
| 10. <i>ionic</i>              | $\text{CdCl}_2$         |
| 11. Sulfur Dioxide            | $\text{SO}_2$           |
| 12. Nitrogen Monoxide         | $\text{NO}$             |
| 13. Sulfur Trioxide           | $\text{SO}_3$           |
| 14. Carbon Tetraiodide        | $\text{CI}_4$           |
| 15. Carbon Tetrachloride      | $\text{CCl}_4$          |
| 16. Phosphorus Trichloride    | $\text{PCl}_3$          |
| 17. Potassium bromide         | <i>ionic</i>            |
| 18. Thallium sulfide          | <i>ionic</i>            |
| 19. nitric acid               | $\text{HNO}_3$          |
| 20. <i>ionic</i>              | $\text{FeI}_2$          |
| 21. Carbon Tetrachloride      | $\text{CCl}_4$          |
| 22. Phosphorus Pentabromide   | $\text{PBr}_5$          |
| 23. Silicon dioxide           | $\text{SiO}_2$          |
| 24. Diarsenic trisulfide      | $\text{As}_2\text{S}_3$ |
| 25. Dihydrogen monoxide/water | $\text{H}_2\text{O}$    |
| 26. Carbon Monoxide           | $\text{CO}$             |
| 27. Dinitrogen tetroxide      | $\text{N}_2\text{O}_4$  |
| 28. Tin (IV) oxide            | <i>ionic</i>            |

# Chemical Names And Formulas Packet Answers

**Marcia L. Gillette, Ira Blei, Wendy  
Gloffke, George Odian**



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enables students with varied backgrounds to proceed together to significant levels of achievement      **Study Guide for Chemical Principles [by] Steven S. Zumdahl** Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples      **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 1982-01-01 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students      Study Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter, 1993      *General, Organic, and Biochemistry Study Guide* Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian's Organic and Biochemistry      Study Guide to Accompany Chemical Principles, Properties, and Reactions Kenneth L. Henold, 1984      Study Guide, Chemical Principles, Fifth Edition, Zumdahl Paul B. Kelter, Steven S. Zumdahl, 2004-04      **Science Spectrum** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03

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