

Name \_\_\_\_\_

*Chemistry Worksheet*  
*Naming & Formula Writing (Ionic)*

**Instructions:** Write the formulas &/or the names for the compounds listed below

- |                          |       |  |       |
|--------------------------|-------|--|-------|
| 1. Sodium nitrate        | _____ | 26. Aluminum chloride                    | _____ |
| 2. Calcium carbonate     | _____ | 27. Iron (III) hydroxide                 | _____ |
| 3. Magnesium oxide       | _____ | 28. Sodium acetate                       | _____ |
| 4. Ammonium sulfide      | _____ | 29. calcium hydroxide                    | _____ |
| 5. Lead (II) sulfate     | _____ | 30. sodium iodate                        | _____ |
| 6. Sodium cyanide        | _____ | 31. Nickel (II) nitrate                  | _____ |
| 7. Potassium hydroxide   | _____ | 32. Iron (II) chloride                   | _____ |
| 8. Silver chloride       | _____ | 33. Magnesium bromide                    | _____ |
| 9. Iron (III) hydroxide  | _____ | 34. Ammonium nitrate                     | _____ |
| 10. Potassium hydroxide  | _____ | 35. Silver bromide                       | _____ |
| 11. Tin (IV) perchlorate | _____ | 36. $\text{Al}(\text{OH})_3$             | _____ |
| 12. Potassium carbonate  | _____ | 37. $\text{NH}_4\text{I}$                | _____ |
| 13. Silver nitrate       | _____ | 38. $\text{Li}_2\text{CO}_3$             | _____ |
| 14. Sodium iodide        | _____ | 39. $\text{CuSO}_4$                      | _____ |
| 15. Ammonium hydroxide   | _____ | 40. KCN                                  | _____ |
| 16. Potassium iodate     | _____ | 41. $\text{Pb}(\text{ClO})_2$            | _____ |
| 17. Lead (IV) oxide      | _____ | 42. BaS                                  | _____ |
| 18. Ammonium hydroxide   | _____ | 43. $\text{ZnSO}_4$                      | _____ |
| 19. Barium sulfate       | _____ | 44. $\text{Pb}(\text{CH}_3\text{COO})_2$ | _____ |
| 20. barium chloride      | _____ | 45. $\text{Ca}(\text{NO}_3)_2$           | _____ |
| 21. Cobalt (II) chloride | _____ | 46. $\text{Fe}_2(\text{CO}_3)_3$         | _____ |
| 22. Sodium carbonate     | _____ | 47. $\text{NH}_4\text{IO}_3$             | _____ |
| 23. Calcium oxide        | _____ | 48. $\text{CaCl}_2$                      | _____ |
| 24. Lead (II) nitrate    | _____ | 49. NaF                                  | _____ |
| 25. Tin (II) chloride    | _____ | 50. $\text{Cu}(\text{NO}_3)_2$           | _____ |

# Chemical Names And Formulas Practice Problems

## Answer Key

**PT Brinkman**



## Chemical Names And Formulas Practice Problems Answer Key:

Addison-Wesley Chemistry ,1990      **Chemistry** John S. Phillips,Cheryl Wistrom,2000      **Educational Technology** ,1982-07      **Prentice Hall Chemistry** ,2000      **The Software Encyclopedia 2001** ,2001      *Chemical Principles* Peter Atkins,Loretta Jones,2007-08 Written for calculus inclusive general chemistry courses *Chemical Principles* helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry s frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of *Chemical Principles* is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding      *Merrill Chemistry* Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998      **Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science** ,2003-11 Prentice Hall *Physical Science Concepts in Action* helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction      *Student Handbook* Southwestern,2005 The *Student Handbook* is designed to provide students with ready access to information with problem solving techniques and study skill guides that enable them to utilize the information in the most efficient manner Amazon com      The Pharmaceutical Era ,1891      Popular Science ,1923-11 *Popular Science* gives our readers the information and tools to improve their technology and their world The core belief that *Popular Science* and our readers share The future is going to be better and science and technology are the driving forces that will help make it better      **Popular Mechanics** ,1936-07 *Popular Mechanics* inspires instructs and influences readers to help them master the modern world Whether it s practical DIY home improvement tips gadgets and digital technology information on the newest cars or the latest breakthroughs in science PM is the ultimate guide to our high tech lifestyle      *Catalog of Copyright Entries. Third Series* Library of Congress. Copyright Office,1965 Includes Part 1 Number 2 Books and Pamphlets Including Serials and Contributions to Periodicals July December      **The Popular Science Monthly** ,1923      *Paperbound Books in Print* ,1970      Mining and Metallurgy ,1926      Popular Mechanics ,1944-08 *Popular Mechanics* inspires instructs and influences readers to help them master the modern world Whether it s practical DIY home improvement tips gadgets and digital technology information on the newest cars or the latest breakthroughs in science PM is

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web hclo 4 is the stronger acid because in a series of oxyacids with similar formulas the higher the electronegativity of the central atom the stronger is the attraction of the central atom for the electrons of the oxygen s the stronger attraction of the oxygen electron results in a stronger attraction of oxygen for the electrons in the o h bond making the hydrogen

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