

Chemical Periodicity:

1. Explain period and group with respect to the periodic table. How many periods and groups are there in the periodic table?

Period = horizontal rows of elements

Groups = Vertical columns of elements

18 groups, 7 periods

2. Explain s, p, d, f block elements with respect to the periodic table.

s block elements (Group I & II) are metallic elements

p block elements (Group III, IV, V, VI, VII) are mainly non metallic elements

d-block and f-block elements are made up of transition elements

3. What are the characteristics of s block elements?

- good conductor of electricity and heat, act as electron donor

4. State the characteristics of the p-block elements. The p-block elements have the outer shell configurations ranging between $s^2 p^1$ to $s^2 p^6$.

- not a good conductor of electricity.

5. State the characteristics of the d-block elements. The d-block elements are also known as the transition elements.

- exhibit various oxidation states

6. State the characteristics of inert gases, why are they inert? The inert or the noble gases are helium, krypton, argon, neon, xenon and radon.

- Because valence shell is fully occupied by electrons and achieve stable configuration

7. Determine the group, period and block for each of the following elements.

A: $1s^2 2s^2$ *Group II, Period 2, s-block elements*

B: $1s^2 2s^2 2p^5$ *Group XVII, Period 2, p-block elements*

C: $1s^2 2s^2 2p^6 3s^2 3p^6 3d^6 4s^2$ *Group VII, Period 4, d-block elements*

D: $1s^2 2s^2 2p^6 3s^2 3p^3$ *Group V, Period 3, P-block elements*

Chemical Periodicity Analysis Question Answers

Nicholas D. Tzimopoulos



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