

Chemistry--Unit 10: Chemical Periodicity  
Practice Problems

I. Classification of the elements

- 1) Use the periodic table to write the electron configuration for silicon.

$1s^2 2s^2 2p^6 3s^2 3p^2$  or abbreviated form\*:  $[\text{Ne}] 3s^2 3p^2$  \*may use this form now!!!

- 2) Use the periodic table to write the electron configuration for iodine.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^5$  or  $[\text{Kr}] 5s^2 4d^{10} 5p^5$

- 3) Which family of elements is characterized by an  $s^2 p^3$  configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:

- one that has 5 outer electrons on the third period of the periodic table **P, phosphorus**
- one with a  $4s^2 4p^5$  electron configuration **Br, bromine**
- the 4<sup>th</sup> period Group 16 element **Se, selenium**

- 5) Identify the elements that have the following outer electron configurations?

- $2s^2 2p^4$  **O, oxygen**
- $4s^2$  **Ca, Calcium**
- $4s^2 3d^{10}$  **Zn, zinc**

- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

- 7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

- 8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d sublevel).

II. Periodic Trends

- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

- 10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

# Chemical Periodicity Practice Problems And Answer Sheet

**Donald A. McQuarrie, Stanley Gill**



## **Chemical Periodicity Practice Problems And Answer Sheet:**

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## **Table of Contents Chemical Periodicity Practice Problems And Answer Sheet**

1. Understanding the eBook Chemical Periodicity Practice Problems And Answer Sheet
  - The Rise of Digital Reading Chemical Periodicity Practice Problems And Answer Sheet
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Periodicity Practice Problems And Answer Sheet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemical Periodicity Practice Problems And Answer Sheet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Periodicity Practice Problems And Answer Sheet
  - Personalized Recommendations
  - Chemical Periodicity Practice Problems And Answer Sheet User Reviews and Ratings
  - Chemical Periodicity Practice Problems And Answer Sheet and Bestseller Lists
5. Accessing Chemical Periodicity Practice Problems And Answer Sheet Free and Paid eBooks
  - Chemical Periodicity Practice Problems And Answer Sheet Public Domain eBooks
  - Chemical Periodicity Practice Problems And Answer Sheet eBook Subscription Services
  - Chemical Periodicity Practice Problems And Answer Sheet Budget-Friendly Options

6. Navigating Chemical Periodicity Practice Problems And Answer Sheet eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemical Periodicity Practice Problems And Answer Sheet Compatibility with Devices
  - Chemical Periodicity Practice Problems And Answer Sheet Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemical Periodicity Practice Problems And Answer Sheet
  - Highlighting and Note-Taking Chemical Periodicity Practice Problems And Answer Sheet
  - Interactive Elements Chemical Periodicity Practice Problems And Answer Sheet
8. Staying Engaged with Chemical Periodicity Practice Problems And Answer Sheet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemical Periodicity Practice Problems And Answer Sheet
9. Balancing eBooks and Physical Books Chemical Periodicity Practice Problems And Answer Sheet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemical Periodicity Practice Problems And Answer Sheet
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemical Periodicity Practice Problems And Answer Sheet
  - Setting Reading Goals Chemical Periodicity Practice Problems And Answer Sheet
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Periodicity Practice Problems And Answer Sheet
  - Fact-Checking eBook Content of Chemical Periodicity Practice Problems And Answer Sheet
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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