

Practice Problems

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SECTION 12.1 THE MOLE: A MEASUREMENT OF MATTER

- Calculate the molar mass of each of the following compounds:
 - phosphorus pentachloride (PCl_5)
 - carbon tetrachloride (CCl_4)
- Calculate the mass percent of each element in the compound:
 - KNO_3
 - Fe_2SO_4
- Calculate the number of moles of each element in 1.00 g of:
 - KNO_3
 - Fe_2SO_4
- Calculate the number of moles of each element in 1.00 g of:
 - KNO_3
 - Fe_2SO_4
- Calculate the mass of 1.00 mole of each of the following compounds:
 - KNO_3
 - Fe_2SO_4
- Calculate the mass of 1.00 mole of each of the following compounds:
 - KNO_3
 - Fe_2SO_4

SECTION 12.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS

- Calculate the mass of each of the following compounds:
 - KNO_3
 - Fe_2SO_4
 - KNO_3
 - Fe_2SO_4
- Calculate the mass of each of the following:
 - 1.00 mole of KNO_3
 - 1.00 mole of Fe_2SO_4
 - 1.00 mole of KNO_3
 - 1.00 mole of Fe_2SO_4
- Calculate the mass of each of the following:
 - 1.00 mole of KNO_3
 - 1.00 mole of Fe_2SO_4
 - 1.00 mole of KNO_3
 - 1.00 mole of Fe_2SO_4
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 - 1.00 mole of KNO_3
 - 1.00 mole of Fe_2SO_4

Chemical Quantities Practice Problems Answer Key

David E. Goldberg



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