

## Balancing Equations and Predicting Products

### 2 Part 1: Balancing Equations

Directions: Balance the following chemical equations and identify the type of reaction.

1.  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  Type of reaction: Synthesis
2.  $\text{NaCl} + \text{KCl} \rightarrow \text{NaCl}$   
Type of reaction: Acid base
3.  $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$   
Type of reaction: Single displacement
4.  $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$   
Type of reaction: Double displacement
5.  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
Type of reaction: Decomposition
6.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$   
Type of reaction: Combustion

### 3 Part 2: Predicting Products

Directions: Given the reactants, predict the products of the following chemical reactions.

1.  $\text{Na} + \text{Cl}_2 \rightarrow ?$
2.  $\text{CaCO}_3 + \text{H}_2\text{SO}_4 \rightarrow ?$
3.  $\text{MgCl}_2 + \text{NaOH} \rightarrow ?$
4.  $\text{HNO}_3 + \text{K} \rightarrow ?$

# Chemical Reactions Balancing And Predicting Products

**Ralph A. Burns**



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