

Chemistry Worksheets Class 10 on Chapter 1 Chemical Reactions and Equations- Set 1

Q-1: Write the balanced chemical equations for each of the following reactions.

- Liquid hydrogen peroxide decomposes to form water and oxygen gas.
- Acetylene gas burns in oxygen to form carbon dioxide and water along with evolution of heat.
- When solid mercury(II) oxide is heated, liquid mercury and oxygen gas are produced.
- Marble dissolves in a hydrochloric acid to give calcium chloride, water and carbon dioxide.
- Copper sulphate on treatment with potassium iodide precipitates cuprous iodide, liberates iodine gas and also forms potassium sulphate.

Q-2: Fill in the blanks

- _____ is the magnetic oxide of iron.
- In reactivity series, manganese lies _____ zinc
- To the reaction between AgNO_3 solution and NaCl solution, the names given to this type of reactions are _____ and _____.
- Proteins decompose to form _____ in our body.
- If copper turnings are added to colourless AgNO_3 solution, it turns _____ after sometime.

Q-3: Which of the following displacement reactions will not take place?

- $\text{Zn(s)} + \text{FeSO}_4\text{(aq)} \rightarrow \text{ZnSO}_4\text{(aq)} + \text{Fe(s)}$
- $\text{Fe(s)} + \text{CuSO}_4\text{(aq)} \rightarrow \text{FeSO}_4\text{(aq)} + \text{Cu(s)}$
- $\text{Cu(s)} + 2\text{AgNO}_3\text{(aq)} \rightarrow \text{Cu(NO}_3)_2\text{(aq)} + 2\text{Ag(s)}$
- $\text{Cu(s)} + \text{FeSO}_4\text{(aq)} \rightarrow \text{CuSO}_4\text{(aq)} + \text{Fe(s)}$

Q-4: Which of the following statements is false?

- Oxidation results in the loss of an electropositive element and the gain of an electronegative element.
- Reduction results in the loss of an electronegative element and the gain of an electropositive element.
- A species that gains electrons is referred to as a reducing agent.
- An oxidising agent is a type of specie that absorbs hydrogen atoms.

Q-5: Why are certain chemicals and medications kept in coloured bottles?

Q-6: Sakshi desired that her home be white washed. She went to the market and bought 20 kg of quick lime, which she dissolved in 60 litres of water. When she added lime to water, she noticed that the water began to boil even though it was not heated. Give an explanation for her observation. Name the product and write the corresponding chemical equation and word equation.

Chemical Reactions Worksheet Chapter 1describing

RJ Alexander



Chemical Reactions Worksheet Chapter 1describing:

This book delves into Chemical Reactions Worksheet Chapter 1describing. Chemical Reactions Worksheet Chapter 1describing is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Chemical Reactions Worksheet Chapter 1describing, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemical Reactions Worksheet Chapter 1describing
 - Chapter 2: Essential Elements of Chemical Reactions Worksheet Chapter 1describing
 - Chapter 3: Chemical Reactions Worksheet Chapter 1describing in Everyday Life
 - Chapter 4: Chemical Reactions Worksheet Chapter 1describing in Specific Contexts
 - Chapter 5: Conclusion
 2. In chapter 1, the author will provide an overview of Chemical Reactions Worksheet Chapter 1describing. This chapter will explore what Chemical Reactions Worksheet Chapter 1describing is, why Chemical Reactions Worksheet Chapter 1describing is vital, and how to effectively learn about Chemical Reactions Worksheet Chapter 1describing.
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 6. In chapter 5, the author will draw a conclusion about Chemical Reactions Worksheet Chapter 1describing. This chapter will summarize the key points that have been discussed throughout the book.
- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Chemical Reactions Worksheet Chapter 1describing.

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