

## Chemical Equations and Reactions

### SECTION 2

**SHORT ANSWER** Answer the following questions in the space provided.

1. Match the equation type on the left to its representation on the right.

\_\_\_ **c** \_\_\_ synthesis



\_\_\_ **d** \_\_\_ decomposition



\_\_\_ **b** \_\_\_ single-displacement



\_\_\_ **a** \_\_\_ double-displacement



2. \_\_\_ **c** \_\_\_ In the reaction described by the equation  $2Al(s) + 3Fe(NO_3)_2(aq) \rightarrow 3Fe(s) + 2Al(NO_3)_3(aq)$ , iron has been replaced by

(a) nitrate.

(c) aluminum.

(b) water.

(d) nitrogen.

3. \_\_\_ **a** \_\_\_ Of the following chemical equations, the only reaction that is both synthesis and combustion is



4. \_\_\_ **b** \_\_\_ Of the following chemical equations, the only reaction that is both combustion and decomposition is



(d)

5. Identify the products when the following substances decompose:

a. **Its separate elements** \_\_\_\_\_ a binary compound

b. **Metal oxide + water** \_\_\_\_\_ most metal hydroxides

c. **Metal oxide + carbon dioxide** \_\_\_\_\_ a metal carbonate

d. **Water + sulfur dioxide** \_\_\_\_\_ the acid  $H_2SO_3$

6. The complete combustion of a hydrocarbon in excess oxygen yields the products

# Chemical Equations And Reactions Chapter 8 Review

**DP Hallahan**



## **Chemical Equations And Reactions Chapter 8 Review:**

**Chemistry (Teacher Guide)** Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study.

**Features**  
Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated, and three-hole punched materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments, dates. There is flexibility in scheduling. Adapt the days to your school schedule.

**Workflow**  
Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade.

**About the Author**  
DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and the Sable Institute of Environmental Studies.

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**Chemistry** Melvin D. Joesten, 1988

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