Equilibrium Constant - Worksheet 1_ANSWERS

1. Write the equilibrium expression for the following reactions:

a.
$$wW + xX = uU + vV$$

$$K = \frac{[W]^{*}[V]^{*}}{[W]^{*}[V]^{*}}$$

b.
$$H_2O_{(g)} + CO_{(g)} \circ H_{2(g)} + CO_{2(g)}$$

$$K = \frac{[H_2][CO_2]}{[H_2O][CO]}$$

c.
$$COCI_{2(g)} = CO_{(g)} + CI_{2(g)}$$

$$K = \frac{[Cl_2][CO]}{[COCl_2]}$$

d.
$$H_{2(g)} + CI_{2(g)} = 2HCI_{(g)} + energy$$

$$K = \frac{[HCI]^2}{[H_2][CI_2]}$$

e.
$$CO_{(g)} + NO_{2(g)} * NO_{(g)} + CO_{2(g)}$$

$$K = \frac{[NO][CO_2]}{[NO_2][CO]}$$

f.
$$Zn_{(n)} + 2Ag^{+}_{(nn)} = Zn^{2+}_{(nn)} + 2Ag_{(n)}$$

$$K = \frac{[Zn^{2^{+}}]}{[Ag^{+}]^{2}}$$

*Note: solids are not included

g.
$$C_2H_{6(g)} = H_{2(g)} + C_2H_{4(g)}$$

$$K = \frac{[H_2][C_2H_4]}{[C_2H_6]}$$

2. Consider the following equilibrium reaction:

$$H_{2(g)} + Br_{2(g)} = 2HBr_{(g)} + energy$$

The K_{eq} for this reaction at 25°C is 1.02. At equilibrium the concentration of HBr is 0.50mol/L. Assuming H_2 and Br_2 are present in equal amounts, calculate the concentration of H_2 at equilibrium.

 $[H_{*}] = 0.50 \text{ mol/L}$

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