

Equilibrium- HW

PSI Chemistry

Name _____

Multiple Choice

Introduction to Equilibrium/General (No calculator)

1) At equilibrium, _____.

- A) all chemical reactions have ceased
- B) the rates of the forward and reverse reactions are equal
- C) the rate constants of the forward and reverse reactions are equal
- D) the value of the equilibrium constant is 1
- E) the limiting reagent has been consumed

2) Which one of the following will change the value of an equilibrium constant?

- A) changing temperature
- B) adding other substances that do not react with any of the species involved in the equilibrium
- C) varying the initial concentrations of reactants
- D) varying the initial concentrations of products
- E) changing the volume of the reaction vessel

3) The equilibrium-constant expression depends on the _____ of the reaction.

- A) stoichiometry
- B) mechanism
- C) stoichiometry and mechanism
- D) the quantities of reactants and products initially present
- E) temperature

4) Consider the following equilibrium:



The equilibrium cannot be established when _____ is/are placed in a 1.0-L container.

- A) 0.25 mol $\text{SO}_2(\text{g})$ and 0.25 mol $\text{O}_2(\text{g})$
- B) 0.75 mol $\text{SO}_2(\text{g})$
- C) 0.25 mol $\text{SO}_2(\text{g})$ and 0.25 mol $\text{SO}_3(\text{g})$
- D) 0.50 mol $\text{O}_2(\text{g})$ and 0.50 mol $\text{SO}_3(\text{g})$
- E) 1.0 mol $\text{SO}_3(\text{g})$

5) Which of the following statements about chemical equilibrium are not true?

- A) The equilibrium constant is dependent on temperature.
- B) The equilibrium constant is independent of the concentrations of reactants or products at certain reaction conditions.
- C) The initial concentrations of reactants and products will affect the equilibrium position.
- D) When starting with only reactants, only the forward reaction will take place until equilibrium is reached.
- E) When starting with only the products, the reverse reaction rate will exceed the forward reaction rate until equilibrium is reached

Chemical Equilibrium Multiple Choice Review

Albert S. Tarendash, Frederick J. Rowe



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