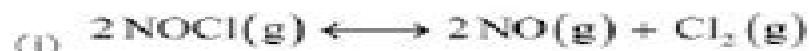


Question 7.4:

Write the expression for the equilibrium constant, K_c for each of the following reactions:



Answer:

$$(i) \quad K_c = \frac{[\text{NO}_{(g)}]^2 [\text{Cl}_{2(g)}]}{[\text{NOCl}_{(g)}]^2}$$

$$(ii) \quad K_c = \frac{[\text{CuO}_{(s)}]^2 [\text{NO}_{2(g)}]^4 [\text{O}_{2(g)}]}{[\text{Cu(NO}_3)_2\text{(s)}]^2}$$

$$= [\text{NO}_{2(g)}]^4 [\text{O}_{2(g)}]$$

$$(iii) \quad K_c = \frac{[\text{CH}_3\text{COOH}_{(aq)}][\text{C}_2\text{H}_5\text{OH}_{(aq)}]}{[\text{CH}_3\text{COOC}_2\text{H}_5\text{(aq)}][\text{H}_2\text{O}_{(l)}]} = \frac{[\text{CH}_3\text{COOH}_{(aq)}][\text{C}_2\text{H}_5\text{OH}_{(aq)}]}{[\text{CH}_3\text{COOC}_2\text{H}_5\text{(aq)}]}$$

$$(iv) \quad K_c = \frac{[\text{Fe(OH)}_{3(s)}]}{[\text{Fe}^{3+}\text{(aq)}][\text{OH}^-\text{(aq)}]^3}$$

$$= \frac{1}{[\text{Fe}^{3+}\text{(aq)}][\text{OH}^-\text{(aq)}]^3}$$

$$(v) \quad K_c = \frac{[\text{IF}_5]^2}{[\text{I}_{2(s)}][\text{F}_2]^5}$$

$$= \frac{[\text{IF}_5]^2}{[\text{F}_2]^5}$$

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