

14

Chemical Periodicity

1. Which of the following orders of ionic radii is correctly represented? [2014]

- (a) $H^+ > H^- > H$ (b) $Na^+ > F^- > O^{2-}$
(c) $F^- > O^{2-} > Na^+$ (d) $Al^{3+} > Mg^{2+} > N^{3-}$

2. Identify the wrong statement in the following. [2012]

- (a) Amongst isoelectronic species, smaller the positive charge on the cation, smaller is the ionic radius
(b) Amongst isoelectronic species, greater the negative charge on the anion, larger is the ionic radius
(c) Atomic radius of the elements increases as one moves down the first group of the periodic table
(d) Atomic radius of the elements decreases as one moves across from left to right in the 2nd period of the periodic table

3. The correct order of the decreasing ionic radii among the following isoelectronic species is [2010]

- (a) $Ca^{2+} > K^+ > S^{2-} > Cl^-$
(b) $Cl^- > S^{2-} > Ca^{2+} > K^+$
(c) $S^{2-} > Cl^- > K^+ > Ca^{2+}$
(d) $K^+ > Ca^{2+} > Cl^- > S^{2-}$

4. Which of the following represents the correct order of increasing electron gain enthalpy with negative sign for the elements O, S, F and Cl? [2010]

- (a) $Cl < F < O < S$ (b) $O < S < F < Cl$
(c) $F < S < O < Cl$ (d) $S < O < Cl < F$

5. Amongst the elements with following electronic configurations, which one may have the highest ionisation energy? [2009]

- (a) $[Ne] 3s^2 3p^2$
(b) $[Ne] 3s^2 3p^3$
(c) $[Ar] 3d^{10}, 4s^2 4p^3$
(d) $[Ne] 3s^2 3p^3$

6. Which one of the elements with the following outer orbital configurations may exhibit the largest number of oxidation states? [2009]

- (a) $3d^3, 4s^2$ (b) $3d^5, 4s^1$ (c) $3d^5, 4s^2$ (d) $3d^2, 4s^2$

7. The correct order of decreasing second ionisation enthalpy of Ti(22), V(23), Cr(24) and Mn(25) is [2008]

- (a) $Cr > Mn > V > Ti$ (b) $V > Mn > Cr > Ti$
(c) $Mn > Cr > Ti > V$ (d) $Ti > V > Cr > Mn$

8. Which of the following electronic configuration of an atom has the lowest ionisation enthalpy? [2007]

- (a) $1s^2, 2s^2 2p^5$ (b) $1s^2, 2s^2 2p^3$
(c) $1s^2, 2s^2 2p^6, 3s^1$ (d) $1s^2, 2s^2 2p^6$

9. Identify the correct order of the size of the following. [2007]

- (a) $Ca^{2+} < K^+ < Ar < S^{2-} < Cl^-$
(b) $Ca^{2+} < K^+ < Ar < Cl^- < S^{2-}$
(c) $Ar < Ca^{2+} < K^+ < Cl^- < S^{2-}$
(d) $Ca^{2+} < Ar < K^+ < Cl^- < S^{2-}$

Chemical Periodicity Answers

Mandell Creighton



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Thermodynamics Chemical Kinetics Electrochemistry Surface Chemistry Metallurgical Operations Chemical Periodicity Hydrogen and its Compounds and s Block Elements p Block Elements Transition Elements d and f Block Elements Coordination Compounds Chemical Analysis General Organic Chemistry Hydrocarbons Alkyl Halides Alcohols Phenols and Ethers Aldehydes And Ketones Carboxylic Acids and their Derivatives Organic Compounds Containing Nitrogen Polymers Biomolecules and Chemistry in Everyday Life Appendix NEET Solved Paper 2018 NEET National Paper 2019 NEET Odisha Paper 2019

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The Principles of chemistry v. 2 Dmitry Ivanovich Mendeleev, 1897

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Table of Contents Chemical Periodicity Answers

1. Understanding the eBook Chemical Periodicity Answers
 - The Rise of Digital Reading Chemical Periodicity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Periodicity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Periodicity Answers
 - Personalized Recommendations
 - Chemical Periodicity Answers User Reviews and Ratings
 - Chemical Periodicity Answers and Bestseller Lists
5. Accessing Chemical Periodicity Answers Free and Paid eBooks
 - Chemical Periodicity Answers Public Domain eBooks
 - Chemical Periodicity Answers eBook Subscription Services
 - Chemical Periodicity Answers Budget-Friendly Options
6. Navigating Chemical Periodicity Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Periodicity Answers Compatibility with Devices
 - Chemical Periodicity Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Periodicity Answers
 - Highlighting and Note-Taking Chemical Periodicity Answers
 - Interactive Elements Chemical Periodicity Answers
8. Staying Engaged with Chemical Periodicity Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Chemical Periodicity Answers
- 9. Balancing eBooks and Physical Books Chemical Periodicity Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Periodicity Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemical Periodicity Answers
 - Setting Reading Goals Chemical Periodicity Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemical Periodicity Answers
 - Fact-Checking eBook Content of Chemical Periodicity Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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