Name: Class:\_ Five Types of Chemical Reaction Worksheet Answers Balance the following reactions and indicate which of the six types of chemical reaction are being 2 NaBr + 1 Ca(OH)<sub>2</sub> → 1 CaBr<sub>2</sub> + 2 NaOH Type of reaction: double displacement að.  $2 NH_1 + 1 H_2SO_4 \rightarrow 1 (NH_4)_2SO_4$ **b**) Type of reaction: synthesis 4 C<sub>5</sub>H<sub>9</sub>O + 27 O<sub>2</sub> → 20 CO<sub>2</sub> + 18 H<sub>2</sub>O Type of reaction: combustion CB. di) 3 Pb + 2 H<sub>2</sub>PO<sub>4</sub> → 3 H<sub>2</sub> + 1 Pb<sub>1</sub>(PO<sub>4</sub>)<sub>2</sub> Type of reaction: single displacement Type of reaction: double displacement et. 1 Li<sub>2</sub>N + 3 NH<sub>4</sub>NO<sub>3</sub> → 3 LiNO<sub>3</sub> + 1 (NH<sub>4</sub>)<sub>3</sub>N Section 3: Predicting the products of chemical reactions 1) 2 Ag + 1 CuSO<sub>4</sub> -> 1 Ag<sub>2</sub>SO<sub>4</sub> + 1 Cu Type: Single Displacement 23 2 Nai + 1 CaCl<sub>2</sub> → 2 NaCl + 1 Cal<sub>2</sub> Type: Double Displacement 30 1 O2 + 1 H2 → 2 H2O Type: Synthesis 400 2 AgNO<sub>2</sub> + 1 BaSO<sub>4</sub> → 1 Ag<sub>2</sub>SO<sub>4</sub> + 1 Ba(NO<sub>2</sub>)<sub>2</sub> Type: Double Displacement 51 2 HCN + 1 CuSO<sub>4</sub> → 1 H<sub>2</sub>SO<sub>4</sub> + 1 Cu(CN)<sub>2</sub> Type: Double Displacement 6) 1 H<sub>2</sub>O + 1 AgI → 1 HI + 1 AgOH Type: Double Displacement 73 4 LiBr + 1 Co(SO<sub>1</sub>)<sub>2</sub> → 2 Li<sub>2</sub>SO<sub>1</sub> + 1 CoBr<sub>4</sub> Type: Double Displacement

Type: Single Displacement

Type: Single Displacement

Type: Single Displacement

Type: Double Displacement

Type: Single Displacement

Type: Double Displacement

Type: Decomposition

Type: Synthesis

Type: Decomposition

Type: Combustion

Type: Combustion

8):

9)

1.0%

1.13

12)

13)

14)

151

1.6)

171

18)

1 LiNO<sub>1</sub> + 1 Ag → 1 AgNO<sub>1</sub> + 1 Li

1 N2 + 2 O2 → 2 NO2

1 H<sub>2</sub>CO<sub>3</sub> → 1 CO<sub>2</sub> + 1 H<sub>2</sub>O

1 AlCl<sub>3</sub> + 3 Cs → 3 CsCl + 1 Al

1 C<sub>0</sub>H<sub>0</sub> + 6 O<sub>2</sub> → 4 CO<sub>2</sub> + 4 H<sub>2</sub>O

1 BaS + 1 PtCl<sub>2</sub> → 1 BaCl<sub>2</sub> + 1 PtS

2 Na₂O → 4 Na + 1 O₂

1 Al(NO<sub>3</sub>)<sub>3</sub> + 1 Ga → 1 Ga(NO<sub>3</sub>)<sub>3</sub> + 1 Al

2 KCl + 1 Mg(OH)₂ → 2 KOH + 1 MgCl₂

1 Zn + 1 Au(NO₂)₂ → 1 Zn(NO₂)₂ + 1 Au

1 CH<sub>3</sub>COOH + 2 O<sub>2</sub> → 2 CO<sub>2</sub> + 2 H<sub>2</sub>O

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