

### ChemQuest 33:

## Limiting Reactants

#### Information: Limiting Reactant

Again, consider the combustion of propane:  $\text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$ . If you had 10 moles of propane to burn, you would need 50 moles of oxygen according to the ratio in the balanced equation. If you only had 10 moles of oxygen you could not combust all 10 moles of propane. The reaction has been limited by the amount of oxygen you have—you don't have enough oxygen to burn all of the propane. In this case, oxygen is called the "limiting reactant" because it limits how much propane can react. Notice that the limiting reactant isn't always the substance that is present in the fewest number of moles. In this example, propane ( $\text{C}_3\text{H}_8$ ) is the "excess reactant" because after the reaction there will be some of it left over. It is important to remember that everything in a chemical equation is related by mole ratios. If you only know the mass (grams) of the substances, you need to convert to moles.

#### Critical Thinking Questions:

- In the above discussion, it was evident that 20 moles of oxygen was not sufficient to combust 10 moles of propane.  
Verify (work out the problem) that if 12.5 moles of propane and 63.2 moles of oxygen were present, then propane is the "limiting reactant", by solving for moles of carbon dioxide.
- Consider the following chemical reaction:  $3 \text{MgCl}_2 + 2 \text{Na}_3\text{PO}_4 \rightarrow 6 \text{NaCl} + \text{Mg}_3(\text{PO}_4)_2$ . Assume that 0.75 mol of  $\text{MgCl}_2$  and 0.65 mol of  $\text{Na}_3\text{PO}_4$  are placed in a reaction vessel. Verify that  $\text{Na}_3\text{PO}_4$  is the excess reactant and  $\text{MgCl}_2$  is the limiting reactant by finding how many moles of sodium chloride could be produced in this reaction.
- Consider the following combustion reaction:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ 
  - If there is 125 g of  $\text{C}_4\text{H}_{10}$  react with 415 g of  $\text{O}_2$ ; what mass of  $\text{H}_2\text{O}$  can be produced?
  - Which substance is the limiting reactant?
- When 412.5 g of calcium carbonate react with 521.9 g of aluminum fluoride, how many grams of aluminum carbonate can be produced?  
$$3 \text{CaCO}_3 + 2 \text{AlF}_3 \rightarrow 3 \text{CaF}_2 + \text{Al}_2(\text{CO}_3)_3$$
- When 277 g of  $\text{Na}_2\text{SO}_4$  reacts with 137 g of  $\text{Ca}_3\text{P}_2$  according to the following balanced equation:  
$$3 \text{Na}_2\text{SO}_4 + \text{Ca}_3\text{P}_2 \rightarrow 2 \text{Na}_3\text{P} + 3 \text{CaSO}_4$$
How many grams of  $\text{Na}_3\text{P}$  are produced? What is the limiting reactant?
- If 312 g of potassium sulfide reacts with 410 g of aluminum phosphide, how many grams of one product (your choice) can be produced? What is the limiting reactant? **\*Write the balanced equation first!**
- If 181.1 g of  $\text{Al}(\text{NO}_3)_3$  react with 102.1 g of  $\text{CaO}$  in a double replacement reaction, how many grams of calcium nitrate will be produced? What is the limiting reactant?  
**\*unbalanced**  
$$\text{Al}(\text{NO}_3)_3 + \text{CaO} \rightarrow \text{Al}_2\text{O}_3 + \text{Ca}(\text{NO}_3)_2$$
- Consider the double replacement reaction between calcium sulfate ( $\text{CaSO}_4$ ) and sodium iodide ( $\text{NaI}$ ). If 34.7 g of calcium sulfate and 58.3 g of sodium iodide are placed in a reaction vessel, how many grams of sodium sulfate are produced? What is the limiting reactant? Write the balanced chemical equation for the reaction and solve.

# Chemical Reactions Limiting Reactants Packet

**Arturo Cuomo**



## **Chemical Reactions Limiting Reactants Packet:**

## Whispering the Techniques of Language: An Mental Quest through **Chemical Reactions Limiting Reactants Packet**

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