

# Lab Manual 11: Introduction to Organic Chemistry

General Chemistry for  
Health Sciences Lab

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# Chemistry Lab Manual For 11

**William Swan Sonnenschein**



## Chemistry Lab Manual For 11:

*Laboratory Manual for Principles of General Chemistry* J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program. The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure, selecting the proper apparatus, employing the proper techniques, systematically analyzing and interpreting the data, and minimizing inherent variables. As a result of good data, a scientific and analytical conclusion is made, which may or may not be right but is certainly consistent with the data. Experiments write textbooks; textbooks don't write experiments. A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered. Further experimentation provides additional cause-effect observations leading to an even better understanding of the experiment. The 11th edition's experiments are informative and challenging while offering a solid foundation for technique, safety, and experimental procedure. The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment. The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit. An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment. Continued emphasis on handling data is supported by the Data Analysis section.

Practical/Laboratory Manual Chemistry Class XI based on NCERT guidelines by Dr. S. C. Rastogi & Er. Meera Goyal Dr. S. C. Rastogi, Er. Meera Goyal, 2020-06-23. An Excellent Book in Accordance with the latest syllabus for Class 11. Prescribed by CBSE, NCERT, and Adopted by Various State Education Boards.

**A. Basic Laboratory Techniques**

1. To cut a glass tube or glass rod
2. To bend the glass rod at an angle
3. To draw a glass jet from a glass tube
4. To bore a cork and fit a glass tube into it

**B. Characterisation and Purification of Chemical Substances**

1. To determine the melting point of the given unknown organic compound and its identification (simple laboratory technique)
2. To determine the boiling point of a given liquid when available in small quantity (simple laboratory method)
3. To prepare crystals of pure potash alum  $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$  from the given impure sample
4. To prepare the pure crystals of copper sulphate from the given crude sample
5. To prepare pure crystals of benzoic acid from a given impure sample

**C. Measurement of pH Values**

1. To determine the pH value of vegetable juices, fruit juices, tap water, and washing soda by using universal pH paper
2. To determine and compare the pH values of solutions of strong acid HCl and weak acid  $CH_3COOH$  of same concentration
3. To study the pH change in the titration of strong base Vs strong acid by using universal indicator paper
4. To study the pH change by common ion  $CH_3COO^-$  ion in case of weak acid  $CH_3COOH$
5. To determine the change in pH value of weak base  $NH_4OH$  in presence of a common ion  $NH_4^+$

**D. Chemical Equilibrium**

1. To study the shift in equilibrium between ferric ions and thiocyanate ions by changing the concentrations of either of the ions
2. To study the shift in equilibrium between  $Co(H_2O)_6^{2+}$  and  $Cl^-$  ions by changing the concentrations of either of the ions

**E. Quantitative Analysis**

1. To

prepare M 10 oxalic acid solution by direct weighing method 2 To prepare M 10 solution of sodium carbonate by direct weighing method 3 To determine the strength of given solution of sodium hydroxide by titrating it against N 10 or M 20 solution of oxalic acid 4 To determine the strength of a given solution of hydrochloric acid by titrating it against a standard N 10 or M 20 sodium carbonate solution F Qualitative Analysis 1 Analysis of Anions 2 Analysis of Cations G Detection of Elements in Organic Compounds 1 To detect the presence of nitrogen sulphur and halogens in a given organic compound by Lassaigne s test 2 To detect the presence of nitrogen sulphur and halogens in the given organic compound sample number by Lassaigne s test INVESTIGATORY PROJECTS A Checking of Bacterial Contamination in Water 1 To check the bacterial contamination in drinking water by testing sulphide ions B Methods of Water Purification 1 To purify water from suspended impurities by using sedimentation 2 To purify water by boiling 3 To purify water by distillation method 4 To purify water by reverse osmosis technique 5 To purify water by GAC method 6 To purify water by bleach treatment 7 To purify water by oxidising agent 8 To purify water by ozone treatment method C Water Analysis 1 To test the hardness of different water samples D Foaming Capacity of Various Soaps 1 To compare the foaming capacity of different washing soaps 2 To study the effect of addition of sodium carbonate on foaming capacity of washing soap E Tea Analysis 1 To study the acidity of different samples of tea leaves tea by using pH paper F Analysis of Fruits and Vegetable Juices 1 To analyse the fruit and vegetable juices for the constituent present in them G Rate of Evaporation 1 To study the rate of evaporation of different liquids H Effect of Acids and Bases on Tensile Strength of Fibres 1 To compare the tensile strength of natural fibres and synthetic fibres 2 To study the effect of acids and bases on tensile strength of different fibres Log Antilog Table General Chemistry I Laboratory Manual, 11E - CHM 111 Brightpoint Community College,2023-05-26 Chemistry Lab Manual Home/Small Christian School Edition (gr. 11) 129551,BJU Press,2005-02-11 The Lab Manual has experiments to demonstrate and reinforce concepts presented in the textbooks All exercises can be performed either in a minimally supplied home or school lab All chemicals used can be purchased locally Grade 11 **General College Chemistry I and II (CHM 11/12)**

**Laboratory Manual** Darrell D. Ebbing,2017 **The Cumulative Book Index** ,1913 **EduGorilla's CBSE Class 11th Chemistry Lab Manual | 2024 Edition | A Well Illustrated, Complete Lab Activity book with Separate FAQs for Viva Voce Examination** , Need an informative and well illustrated Lab Manual CBSE Class 11th Chemistry Lab Manual is here for you The Lab Manual provides comprehensive steps for guiding students through each experiment Rigorously researched content prepared by a team of educators writers editors and proofreaders CBSE Class XI Chemistry Lab Manual has properly labeled high resolution diagrams and graphs A separate section on Viva Questions has been included to aid students in their Viva examination The Lab Manual explains the complex topics through detailed illustrations and lucid language making them simple to grasp Worksheets have been provided in CBSE Class 11th Chemistry Lab Manual for doing rough work

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