

Topic One –Answer Key

Part A

A glass bulb of volume 189mL was filled with HCl gas to a pressure of 108 torr at 25°C. How many moles of gas were in the bulb?

Do you understand all of the terms?

Volume, pressure, torr, moles

What is the question asking for?

Moles of HCl gas

Write down your logic for solving this problem. What are the steps to solve this problem?

Equation: $PV=nRT$

Given: P, V, and T

R is a constant. When I looked it up I found

$$R = 8.314 \text{ J/(mol K)} = 0.08206 \text{ (L atm)/(mol K)}$$

Pick one of the constants

Want to find "n"

Convert P, V, and T to match the units in the constant

Plug everything into the equation

Answer the question.

$$V = 189\text{mL}$$

$$\text{HCl } p = 108 \text{ torr}$$

$$T = 25^\circ\text{C}$$

$$\text{Moles} = ?$$

$$PV=nRT$$

$$R = 0.08206 \text{ (L atm)/(mol K)}$$

$$1 \text{ atm} = 760 \text{ torr}$$

$$\text{K} = ^\circ\text{C} + 273.15^\circ\text{C}$$

$$\frac{PV}{RT} = n$$

$$\left(\frac{108\text{torr}}{1} \times \frac{1\text{atm}}{760\text{torr}} \right) \left(\frac{189\text{mL}}{1} \times \frac{1\text{L}}{10^3\text{mL}} \right) = X\text{moles} = 0.001098 \text{ moles}$$
$$\left(273.15^\circ\text{C} + 25^\circ\text{C} \right) \left(0.08206 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}} \right)$$

What concepts do you have to know to answer the questions?

(The chemistry question above is "algorithmic", meaning you have to recall an equation and plug in information to solve.)

Nature of a gas

Ideal gas law

Unit conversions (including the temperature equations)

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