

Equations and Values Related to the 2002 AP® Exam

Electron in terms of m_e $\lambda = 2.82 \times 10^{-10} \text{ m}$ $\frac{h}{mv} = \frac{h}{m\lambda f}$ $\frac{h}{m\lambda} = \frac{h}{m\lambda} \cdot \frac{1}{\lambda} = \frac{h}{m\lambda^2}$ $\frac{h}{m\lambda^2} = \frac{h}{m\lambda^2} \cdot \frac{1}{\lambda} = \frac{h}{m\lambda^3}$

$E = hf$ $E = hf = \frac{hc}{\lambda}$ $E = hf = \frac{hc}{\lambda} \cdot \frac{1}{\lambda} = \frac{hc}{\lambda^2}$

$E = hf = \frac{hc}{\lambda} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{1.21 \times 10^{-6} \text{ m}} = 5.48 \times 10^{-19} \text{ J}$ $E = hf = \frac{hc}{\lambda} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{1.21 \times 10^{-6} \text{ m}} = 5.48 \times 10^{-19} \text{ J}$ $E = hf = \frac{hc}{\lambda} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}{1.21 \times 10^{-6} \text{ m}} = 5.48 \times 10^{-19} \text{ J}$

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