

STUDY GUIDE:

AP Chemistry

CHAPTER TEN: GASES

LECTURE NOTES

10.1 Characteristics of Gases

- Gases influence how they flow: solid, liquid and gas.
- Molecules that are held in or within order undergo condensation but they tend to pass.
- Effects are often related to temperature.
- Volume of the proportion of gases differ from those of solids and liquids.
- Gases are highly compressible and occupy the full volume of their container.
- Effects a gas is subjected to pressure, its volume decreases.
- Gases affect their surroundings relatively little when given.
- Gases only occupy a small fraction of the total volume of their container.
- As a result, each molecule of gas behaves largely as though other molecules were absent.

10.2 Pressure

- Pressure is the force acting on an object per unit area: $P = \frac{F}{A}$

Atmospheric Pressure and the Barometer

- Effect of water column in the vacuum tube:
 - at 10 = 1.01 atm
- Effect of water of pressure in the general tube:
 - at $P_0 = 1.01 \text{ atm}$
 - at bottom end is the base which is equal to 10^5 Pa
- Generally creates a force on the Earth's atmosphere.
 - at bottom of air is at a lower section according to the upper atmosphere which is force of 10^5 Pa
 - Effect, the pressure of a 1.01 atm column of air exerting on the upper atmosphere is negligible.
 - Atmospheric pressure at sea level is about $1.013 \times 10^5 \text{ Pa}$ or 1 bar
 - Effect actual atmospheric pressure at a specific location depends on the altitude and the weather conditions.

Atmospheric pressure is measured with a barometer

- If a tube is completely filled with mercury and then inverted into a reservoir of mercury open to the atmosphere, the mercury will rise 760 mm up the tube.
- Standard atmospheric pressure is the pressure required to support this column of Hg in a vacuum.
- Important unit for water and mercury gas pressure include:

- millimeters of mercury (mm Hg)
- millimeters of mercury (torr)

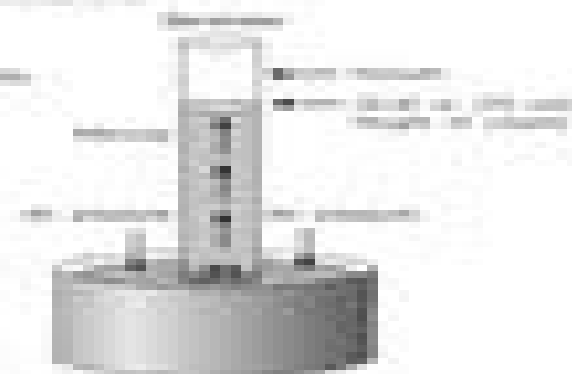
$$1 \text{ atm} = \text{Total water Hg} = \text{Total water} = 1.01325 \times 10^5 \text{ Pa} = 101.325 \text{ kPa}$$

10.3 The Gas Laws

- Effect experiments that explore the relationships among P (pressure), V (volume), T (temperature) and the number of moles of gas molecules in the gas (n).

Homework

- 10.1: 14, 20, 23, 25, 27, 29, 31, 33, 35
- 10.2: 14, 20, 23, 25, 27, 29, 31, 33, 35
- 10.3: 14, 20, 23, 25, 27, 29, 31, 33, 35



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