

STUDY GUIDE:

AP Chemistry

CHAPTER TEN: GASES

LECTURE NOTES

10.1 Characteristics of Gases

- Gases influence how they flow: solid, liquid and gas.
- Molecules that are held in or with other molecules condense but they can't go past.
- Gases are able to expand to fill space.
- Volume of the proportion of gases differ from those of solids and liquids.
- Gases are highly compressible and occupy the full volume of their container.
- Gases are subject to pressure, to volume changes.
- Gases affect their surroundings relatively little when given.
- Gases only occupy a small fraction of the total volume of their container.
- As a result, the behavior of gas behavior largely is thought with independent molecules.

10.2 Pressure

- Pressure is the force acting on an object per unit area: $P = \frac{F}{A}$

Atmospheric Pressure and the Barometer

- The 10 units of force in the vacuum (Hg):
 - 10.13 in Hg
- The 10 units of pressure in the general (Pa):
 - 101.3 kPa
 - 1.013 bar
 - 1.013 atm
- Gravity creates a force on the Earth's atmosphere.
- Columns of air 1 m² at lower altitude according to the upper atmosphere create a force of 101.3 kPa.
- Effect, the pressure of a 1 m² column of air pressing on the lower atmosphere is 101.3 kPa.
- Atmospheric pressure at sea level is about 101.3 kPa or 1 bar.
- Effect, actual atmospheric pressure at a specific location depends on the altitude and the weather conditions.

Atmospheric pressure is measured with a barometer

- If a tube is completely filled with mercury and then inverted into a reservoir of mercury open to the atmosphere, the mercury will rise 760 mm up the tube.
- Standard atmospheric pressure is the pressure required to support that column of Hg in a vacuum.
- Important unit for units used to express gas pressure include:

- millimeters of mercury (mm Hg)
- millimeters of mercury (torr)

$$1 \text{ atm} = 760 \text{ mm Hg} = 760 \text{ torr} = 1.01325 \times 10^5 \text{ Pa} = 101.325 \text{ kPa}$$

10.3 The Gas Laws

- The equations that express the relationships among P, temperature, V, pressure, V, volume and the number of moles of gas are listed in the gas laws.

Homework

10.1 Pa, 10.1 kPa, 10.1 Pa, 10.1 Pa, 10.1 Pa

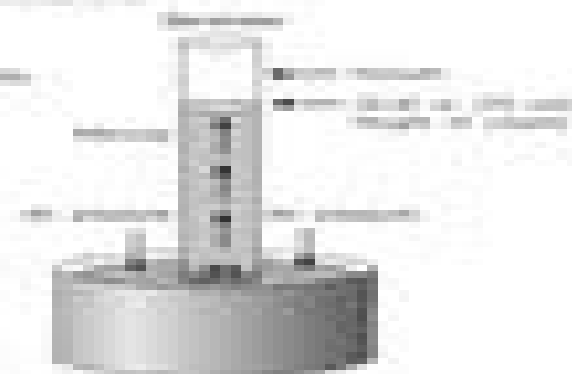
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 - The Rise of Digital Reading Chemistry 10study Guides
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 10study Guides
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 10study Guides
 - User-Friendly Interface
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 - Personalized Recommendations
 - Chemistry 10study Guides User Reviews and Ratings
 - Chemistry 10study Guides and Bestseller Lists
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- Chemistry 10study Guides Enhanced eBook Features
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 - Interactive Elements Chemistry 10study Guides
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 10study Guides
- 9. Balancing eBooks and Physical Books Chemistry 10study Guides
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 10study Guides
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 10study Guides
 - Setting Reading Goals Chemistry 10study Guides
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 10study Guides
 - Fact-Checking eBook Content of Chemistry 10study Guides
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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