



$$(1.98 \text{ g H}_2 / 7.575 \text{ g H}_2) \times 100\% = \mathbf{26.1\%}$$

- b. Suppose your % yield was 68.3% instead, how much H₂ would you make?
(or 7.575 g H₂) × 100% = 68.3%

$$x = \mathbf{5.19 \text{ g H}_2}$$



- a. How many grams of iron can be made from 118 g of Fe₂O₃ and 12.7 g C?



- b. What is the limiting reagent?



- c. After the above reaction, you produce 25.6 g of Fe. What is your % yield?
(25.6 g Fe / 63.9 g Fe) × 100% = $\mathbf{55.7\%}$



- a. How many grams of sugar can be made from 50.0 g of CO₂ and 50.0 g of water?



- b. What is the limiting reagent?



- c. What is the excess reagent?



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