

Question 1. Calculate the molecular mass of the following:

(i) H_2O (ii) CO_2 (iii) CH_4

Answer: (i) Molecular mass of $\text{H}_2\text{O} = 2(1.008 \text{ amu}) + 16.00 \text{ amu} = 18.016 \text{ amu}$

(ii) Molecular mass of $\text{CO}_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$

(iii) Molecular mass of $\text{CH}_4 = 12.01 \text{ amu} + 4 (1.008 \text{ amu}) = 16.042 \text{ amu}$

Question 2. Calculate the mass percent of different elements present in sodium sulphate ($\text{Na}_2 \text{SO}_4$).

Answer:

$$\text{Mass \% of an element} = \frac{\text{Mass of that element in the compound}}{\text{Molar mass of the compound}} \times 100$$

Now, $\text{Molar mass of } \text{Na}_2\text{SO}_4 = 2 (23.0) + 32.0 + 4 \times 16.0$
 $= 142 \text{ g mol}^{-1},$

$$\begin{aligned}\text{Mass percent of sodium} &= \frac{46}{142} \times 100 \\ &= 32.39 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of sulphur} &= \frac{32}{142} \times 100 \\ &= 22.54 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percent of oxygen} &= \frac{64}{142} \times 100 \\ &= 45.07 \%\end{aligned}$$

Question 3. Determine the empirical formula of an oxide of Iron which has 69.9 % iron and 30.1 % dioxygen by mass.

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Clifford Lane



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