Grade 12 Chemistry Exam Review

Below is a list of major topics covered on the exam for this course. If a topic is not listed, it will NOT be on the exam.

Your studying should begin with your class notes, assigned problems, textbook problems and extra review questions suggested below. This list is a guide and some specific questions on the exam may not appear here.
Good luck ☺

Unit #1 — Energy Changes Text Reference: Chapter 5

- ☐ Energy changes in reactions (system, surroundings, endothermic vs. exothermic)
- ☐ Specific Heat Capacity
- ☐ Calorimetry, $q=mc\Delta T$
- ☐ Molar enthalpies
- ☐ Representing Enthalpy Changes (Thermochemical Equations, Potential Energy Diagrams, etc.)
- ☐ Hess' Law
- ☐ Standard Enthalpies of Formation

Sample Exam Questions

1. Given the following information:

$2CH_3OH_{(l)} + 3O_{2(g)} \rightarrow 2CO_{2(g)} + 4H_2O_{(l)}$, $\Delta H^\circ = -727\ kJ \cdot mol^{-1}$

$\Delta H_f^\circ\ CO_{2(g)} = -394\ kJ \cdot mol^{-1}$, $\Delta H_f^\circ\ H_2O_{(l)} = -286\ kJ \cdot mol^{-1}$

What is the standard heat formation, ΔH_f° ($kJ \cdot mol^{-1}$) of $CH_3OH_{(l)}$?

A. -602 B. -1329 C. 1329 D. 602

2. When 789 g of liquid ammonia is cooled from 82.7 °C to 25.0 °C it loses 214 kJ of heat. What is the specific heat capacity of liquid ammonia?

A. 0.213 B. 4.19 C. 0.0047 D. 2.06 E. 4.70

3. Identify the correct expressions that can be used to determine the change in thermal energy for the following reaction:

$C_4H_{10(g)} + \frac{13}{2}O_{2(g)} \rightarrow 4CO_{2(g)} + 5H_2O_{(l)} + heat$?

A. $\Delta H^\circ_r = [(4\ mol)(\Delta H^\circ_f, CO_{2(g)}) + (5\ mol)(\Delta H^\circ_f, H_2O_{(l)})] + [(1\ mol)(\Delta H^\circ_f, C_4H_{10(g)}) + (\frac{13}{2}\ mol)(\Delta H^\circ_f, O_{2(g)})]$ B. $\Delta H^\circ_r = [(1\ mol)(\Delta H^\circ_f, C_4H_{10(g)}) + (\frac{13}{2}\ mol)(\Delta H^\circ_f, O_{2(g)})] - [(4\ mol)(\Delta H^\circ_f, CO_{2(g)}) + (5\ mol)(\Delta H^\circ_f, H_2O_{(l)})]$

C. $\Delta H^\circ_r = [(1\ mol)(\Delta H^\circ_f, C_4H_{10(g)}) + (\frac{13}{2}\ mol)(\Delta H^\circ_f, O_{2(g)})] - [(4\ mol)(\Delta H^\circ_f, CO_{2(g)}) + (5\ mol)(\Delta H^\circ_f, H_2O_{(l)})]$

D. $\Delta H^\circ_r = [(4\ mol)(\Delta H^\circ_f, CO_{2(g)}) + (5\ mol)(\Delta H^\circ_f, H_2O_{(l)})] - [(1\ mol)(\Delta H^\circ_f, C_4H_{10(g)}) + (\frac{13}{2}\ mol)(\Delta H^\circ_f, O_{2(g)})]$

E. $\Delta H^\circ_r = [(1\ mol)(\Delta H^\circ_f, CO_{2(g)}) + (\frac{4}{5}\ mol)(\Delta H^\circ_f, H_2O_{(l)})] - [(1\ mol)(\Delta H^\circ_f, C_4H_{10(g)}) + (\frac{13}{2}\ mol)(\Delta H^\circ_f, O_{2(g)})]$

4. For which of the following statements is correct for the reaction

$2HO2(g) \rightarrow H_{2(g)} + 2O_{2(g)} + 66.4\ kJ$

A. exothermic, potential energy greater in reactants
B. exothermic, potential energy greater in products
C. endothermic, potential energy greater in products
D. endothermic, potential energy less in reactants
E. endothermic, potential energy less in products

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