

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

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1. Understanding the eBook Chemistry 101 Final Exam Study Guide
 - The Rise of Digital Reading Chemistry 101 Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 101 Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 101 Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 101 Final Exam Study Guide
 - Personalized Recommendations
 - Chemistry 101 Final Exam Study Guide User Reviews and Ratings
 - Chemistry 101 Final Exam Study Guide and Bestseller Lists
5. Accessing Chemistry 101 Final Exam Study Guide Free and Paid eBooks
 - Chemistry 101 Final Exam Study Guide Public Domain eBooks
 - Chemistry 101 Final Exam Study Guide eBook Subscription Services
 - Chemistry 101 Final Exam Study Guide Budget-Friendly Options

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 - Chemistry 101 Final Exam Study Guide Compatibility with Devices
 - Chemistry 101 Final Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Chemistry 101 Final Exam Study Guide
 - Interactive Elements Chemistry 101 Final Exam Study Guide
8. Staying Engaged with Chemistry 101 Final Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 101 Final Exam Study Guide
9. Balancing eBooks and Physical Books Chemistry 101 Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 101 Final Exam Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry 101 Final Exam Study Guide
 - Setting Reading Goals Chemistry 101 Final Exam Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry 101 Final Exam Study Guide
 - Fact-Checking eBook Content of Chemistry 101 Final Exam Study Guide
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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