

**CHEM 110**  
**Extra In-Class Practice Problems**  
**Lecture Set 3**

1. Answer the following questions dealing with Avogadro's number.

A. How many molecules present in 2.0 moles  $\text{CH}_4$ ?

$$\frac{2.0 \text{ mol } \text{CH}_4}{1 \text{ mol } \text{CH}_4} \times \frac{6.022 \times 10^{23} \text{ molecules}}{1 \text{ mol } \text{CH}_4} = 1.2 \times 10^{24} \text{ molecules } \text{CH}_4$$

How many atoms are present?

$$\frac{1.2 \times 10^{24} \text{ molecules } \text{CH}_4}{1 \text{ molecule } \text{CH}_4} \times \frac{5 \text{ atoms}}{1 \text{ molecule } \text{CH}_4} = 6.0 \times 10^{24} \text{ atoms}$$

B. How many formula units of  $\text{CaCl}_2$  are present in 0.543 mol  $\text{CaCl}_2$ ?

$$\frac{0.543 \text{ mol } \text{CaCl}_2}{1 \text{ mol } \text{CaCl}_2} \times \frac{6.022 \times 10^{23} \text{ f.u.}}{1 \text{ mol } \text{CaCl}_2} = 3.27 \times 10^{23} \text{ f.u. } \text{CaCl}_2$$

C.  $9.81 \times 10^{21}$  molecules  $\text{O}_2$  is how many moles?

$$\frac{9.81 \times 10^{21} \text{ molecules } \text{O}_2}{6.022 \times 10^{23} \text{ molecules}} \times \frac{1 \text{ mol } \text{O}_2}{6.022 \times 10^{23} \text{ molecules}} = 0.0163 \text{ mol } \text{O}_2$$

D. How many atoms are present in 0.371 mol  $\text{P}_2\text{O}_5$ ?

$$\frac{0.371 \text{ mol } \text{P}_2\text{O}_5}{1 \text{ mol } \text{P}_2\text{O}_5} \times \frac{6.022 \times 10^{23} \text{ molecules } \text{P}_2\text{O}_5}{1 \text{ mol } \text{P}_2\text{O}_5} \times \frac{7 \text{ atoms}}{1 \text{ molecule } \text{P}_2\text{O}_5} = 1.56 \times 10^{24} \text{ atoms}$$

# Chemistry 110 Practice Problems

**University of California (1868-1952)**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the text 'University of California (1868-1952)'.

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