Solubility Rules Practice Worksheet

Classify each substance as being soluble or insoluble in water.

Mg(PO₄)₂ –

II. KOH – _____

III. NiCl₂ -

IV. NH4OH - _____

V. Hg.SO₄ –

VI. KBr -

VII. Pb(CO,) - _____

VIII. Pbl, - _____

IX. BaSO, - _____

X. NiCl, -

2. Show the ions that formed the following compounds:

I. Zn₂(PO₄)₂

II. AlbSa

III. Iron (III) sulfide

IV. Ammonium cyanide

Form 4 water-soluble compounds by combining ions from the ions below:

CI

CO.2 PO.3

1.70

Serv

Identify the precipitate in the following reaction. Circle the correct answer.

a. Li,CO, + Co(CH,COO), → 2 LiCH,COO + CoCO,

b. Pb(NO,), + Li,SO, → PbSO, + 2 LiNO,

Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
Zn ₃ (PO ₄) ₂		
	Silver bromide	
KNO ₃		
	Aluminum sulfide	
	Silver acetate	

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