# Chemistry 30 - Formulas and Information

# Thermochemistry:

Formula:	Variables:	When to use:
Q = mc∆t	Q = total heat/enthalpy (J or kJ) m = nouse (g) c = specific heat capacity (J/g-C or kJ/g-C) ΔI = temperature change (+C)	If given the specific heat capacity, mass and temperature change of one substance.
n = modes of substance reacting (n = cv or mM) (mod)  H = moder best/ enthalpy of substance reacting (J/mod or kJ/mod)  m = mass of water (tg = 1 mL)  c = specific best capacity of water (J/g-C or kJ/g-C)  Δt = term, change of water (-C)		Calconnetry questions - if given information about a reacting substance and water (remarker water is present to determine the energy released or abouthed by the reaction) - usually asked to find the molar enthalpy (H)
$\Delta H = \sum_{i} n_i H(\text{products})$ $= \sum_{i} n_i H(\text{products}) = \sum_{i} n_i H$		Hese's Law: - if only have a reaction with no other energy information - remember: - use data bookdet for heats of formations of several compounds - elements have a heat of formation of pero
% efficiency =    % efficiency (%)   mergy output (J or kJ released by a system)   mergy input (J or kJ absorbed by a system)   mergy input (J or kJ absorbed by a system)		If asked for % efficiency and given the energy released and used by a system

# Electrochemistry:

Formula:	Variables:	When to use:
$E_{ot}$ = $E_{od}$ + $E_{ot}$	E'_= net voltage of a cell (V) E'_= voltage of the reduction half reaction (V) E'_= voltage of the undation half reaction (V)	If a fixed to find the voltage of a cell  the voltages come fixes the data booklet  the condution half reaction comes from flipping the reduction half reaction in the data booklet (remember to flip the voltage sign as well, but you don't need to multiply the voltage if reaction was multiplied for equal electron transfers)
q = It	q = charge (C) I = current (C/t or I) I = time (i)	Electrical Doublemetry:  - if given info about an electrolytic process such as running a current and the length of time.
n <sub>e</sub> = <u>q</u> or <u>lt</u> F F	n_= moles of electrons (mol) q = charge (C) l = current (C/r or l) t = time (r) F = 9.65 x 10° C-tnol (Farndag's Constant) n = moles of reacting substance (mol)	Electrical Stoichiometry:  - if given info about an electrolytic process such as running a current and the length of time used to produce a solid metal  - need to write the half reaction of the metal (substance) being electroplated or refined  - use the n, to calculate the n of the reacting substance (multiply n by the coefficient ratios wanted/given)

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