

CHEM 1046 Review Sheet for Test 3 (study notes & quizzes in addition to review sheet)

1. Carbon tetrachloride reacts at high temperatures with oxygen to produce two toxic gases, phosgene and chlorine.



Calculate K_c for the reaction $2\text{CCl}_4(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{COCl}_2(\text{g}) + 2\text{Cl}_2(\text{g})$.

- a) 4.4×10^2
- b) 8.8×10^2
- c) 1.9×10^{-2}
- d) 1.9×10^{-2}**
- e) 2.3×10^{-2}

Answer:

D

2. What is the value of ΔG° for the following reaction if ΔG°_f for $\text{H}_2\text{O}(\text{g}) = -228.57 \text{ kJ/mol}$, and ΔG°_f for $\text{CO}_2(\text{g}) = -394.36 \text{ kJ/mol}$?



- A) 81
- B) 8.5
- C) 2.2
- D) -890**
- E) 890

Answer: D

3. A container was charged with hydrogen, nitrogen, and ammonia gases at 120°C and the system was allowed to reach equilibrium. What will happen if the volume of the container is increased at constant temperature?



- a) There will be no effect.
- b) More ammonia will be produced at the expense of hydrogen and nitrogen.
- c) Hydrogen and nitrogen will be produced at the expense of ammonia.**
- d) The equilibrium constant will increase.
- e) The equilibrium constant will decrease.

Ans

wer: C

4. What is conjugate acid of NH_3 ?

- a) NH_4^+

Chemistry 6 7 Test Answers

Hussin A.Rothana



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