

SEMESTER 2 STUDY GUIDE

You must show your work for any question that requires you to do calculations

Unit 8 and Unit 9 - The Mole and Stoichiometry

1. Give a definition of the term mole
a specific number of particles, which is approximately 6.022×10^{23} particles, also known as Avogadro's number.
2. How many liters of a gas are in one mole?
one mole of any ideal gas occupies 22.4 liters.
3. How many particles/atoms/molecules/formula units are in one mole?
 6.022×10^{23} particles.
4. How do you determine one mole when it comes to the mass of a substance?
(g/mol)
5. Convert 2 moles of O_2 gas into mass in grams.
32 g/mol. Therefore, 2 moles of O_2 gas is equal to 64 grams of O_2 gas.
6. Convert 3.5 moles of NaCl into formula units.
 6.022×10^{23} particles/mol. Therefore, 3.5 moles of NaCl is equal to 2.11×10^{24} formula units of NaCl.
7. Convert 0.25 moles of chlorine gas into liters.
 $V = nRT/P = (0.25 \text{ mol})(0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K})(273 \text{ K})/(1 \text{ atm}) = 5.57$ liters of chlorine gas.
8. Convert 56.25 grams of water into moles.
18 g/mol. Therefore, 56.25 grams of water is equal to 3.125 moles of water.
9. Convert 15.65 liters of N_2 gas into moles.
 $n = PV/RT = (1 \text{ atm})(15.65 \text{ L})/(0.0821 \text{ L}\cdot\text{atm/mol}\cdot\text{K})(273 \text{ K}) = 0.680$ moles of N_2 gas.
10. Convert 1.65×10^{24} atoms of Mg into moles.
 6.022×10^{23} particles/mol. Therefore, 1.65×10^{24} atoms of Mg is equal to 27.4 moles of Mg.
11. Convert 1.65×10^{24} atoms of Mg to mass.
24.31 g/mol. Therefore, 1.65×10^{24} atoms of Mg is equal to 667 grams of Mg.
12. Convert 6.25 grams of CO_2 to molecules of CO_2
 $(0.142 \text{ mol})(6.022 \times 10^{23} \text{ molecules/mol}) = 8.56 \times 10^{22}$ molecules of CO_2 .
13. Determine the limiting reactant if you have a 3.86 grams sample of NO_2 and a 45.25 gram sample of O_2 .
 $2NO_2(g) + O_2 \rightarrow 2NO$
 $(3.86 \text{ g})/(46.01 \text{ g/mol}) = 0.084 \text{ mol of } NO_2$
14. What is the maximum amount of NO that can be produced from the reaction in #13?
 $2NO_2(g) + O_2(g) \rightarrow 2NO(g)$

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James Millikin University



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