

CHEM 101 Worksheet, Chapter 2: Atoms and Molecules ANSWERS

Remember: $\text{amu} = \#p^+ + \#n$, so $\#n = \text{amu} - \#p^+$. For neutral elements, $\#p^+ = \#e^-$.

Cations (+) lose electrons, anions (-) gain electrons

1. Phosphorus (P) is element # 15, $Z = 15$, 31 amu. It has ___ protons, ___ neutrons, ___ electrons
C. 15, 16, 15

2. An isotope of the element bromine, Br, has 35 protons, it has 80 atomic mass units.

$Z =$ ___, ___ neutrons, ___ electrons

E. 35, 45, 35

3. Sodium, Na is element #11 The $^{23}\text{Na}^+$ ion has ___ protons, ___ neutrons, ___ electrons

E. 11, 12, 10 Na⁺ is a +1 cation, so it has lost one electron!

4. Chlorine is element #17. The chloride ion, Cl^- , of isotope ^{35}Cl has ___ p^+ , ___ n , ___ e^-

D. 17, 18, 18 chloride is a -1 anion, so it has one added electron!

5. Sulfur is element #16, 32 amu. The sulfide ion $^{32}\text{S}^{2-}$ has ___ p^+ , ___ n , ___ e^-

B. 16, 16, 18 The sulfide anion has 2 negative charges, so 2 extra electrons!

The empirical formula is derived from the molecular formula by dividing by the largest common denominator. The empirical formula represents the ratio of the atoms in a molecule.

6. D. C_4H_8 For octane, C_8H_{18} the highest number we can divide by is 2, so the empirical formula is C_4H_8 .

7. A. CH_2 The ratio of carbon to hydrogen atoms in propene, C_3H_4 is 1:2

In the tables below, fill in the correct name or formula

8. Give formulas for the following compounds.

Example: sodium chloride NaCl

a. potassium carbonate K_2CO_3 the carbonate anion, CO_3^{2-}

b. copper(II) sulfide CuS it is copper(II), so Cu^{2+}

c. iron(III) sulfite $\text{Fe}_2(\text{SO}_3)_3$ note the brackets, sulfite is SO_3^{2-} from sulfurous acid, H_2SO_3

d. sodium dihydrogen phosphate NaH_2PO_4 the phosphate anion = PO_4^{3-}

e. nitric acid HNO_3 Know the common "mineral acids", HCl, HF, HBr, HI, HNO_3 , H_2SO_4 , H_3PO_4 !!

f. phosphoric acid H_3PO_4

g. aluminum nitrite $\text{Al}(\text{NO}_2)_3$ Al is in 3rd group, so Al^{3+} . nitrite is the anion from nitrous acid, HNO_2

h. dinitrogen tetroxide N_2O_4 a molecular compound

i. carbon monoxide CO carbon dioxide is CO_2

j. magnesium sulfate heptahydrate $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$ Mg in the 2nd group, so Mg^{2+}

k. sulfuric acid H_2SO_4 l. sulfurous acid H_2SO_3

m. hypochlorous acid HClO n. chloric acid HClO_3 o. periodic acid HIO_4 , think of perchloric acid HClO_4

9. Name the following compounds

Example: KCl potassium chloride

a. K_2SO_4 potassium sulfate

b. SO_3 sulfur trioxide

c. Fe_2O_3 iron(III)oxide We know it is Fe(III) because oxide = O^{2-}

d. $\text{Ba}(\text{OH})_2$ barium hydroxide

e. $\text{Al}_2(\text{CO}_3)_3$ aluminum carbonate

f. HClO_3 chloric acid

g. CuClO_3 copper(I)chlorate we know it must be Cu^+ because chlorate from HClO_3 is ClO_3^-

h. HClO_4 perchloric acid

i. NaClO sodium hypochlorite

j. $\text{Mg}(\text{NO}_3)_2$ magnesium nitrate

Chem 2014 Response Answers

Mina Lebitz



Chem 2014 Response Answers:

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identifying mechanism of action their potential use as biomarkers and few of them investigate a therapeutic usage However there are some issues to be achieved on the path to their clinical application This research topic offers a common place to discuss current and novel clinical applications of EVs pointing on future directions We encouraged the submission of original articles reviews hypothesis controversies future perspectives and personal viewpoints on the following topics of interest but not limited to Contribution of EVs to better understand the pathology of immunological diseases Standardization of isolation and quantification protocols in the daily clinical practice Possible applications of EVs as clinical biomarkers diagnostic prognostic and evolution marker Therapeutic role of EVs being vehicles of specific cargo current clinical trials Novel immunological functions of EVs

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