

Experiment 1 - Volumetric Analysis: Unknown Al^{3+} Concentration

- Given an aqueous solution of $\text{p}(\text{HCl})_3$ of unknown concentration
- Goal is determine the concentration of Al^{3+}
- Chemistry used: solubility
- Method: convert Al^{3+} (soluble, clear) to $\text{Al}(\text{OH})_3$ (insoluble, yellow)
- pH
 - Low pH ($\text{pH} \approx 1$) test across $\text{H}^+ + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+$
 - High pH convert protein stain $\rightarrow$ by test *insoluble*
 - Very high pH: $\text{Al}^{3+} + \text{OH}^- \rightarrow [\text{Al}(\text{OH})_4]^-$
 - Neutral pH: Al^{3+} (soluble)
 - pH 5: buffer + ammonium acetate ($\text{pH}_a/\text{pH}_b/\text{pH}_c$)
- Calculations:
 - (mass of oxalate + precipitate) - (mass of oxalate) = mass of precipitate
 - (mass of precipitate)/(molar mass) $\rightarrow$ mol of $\text{Al}(\text{OH})_3$ = mol of Al^{3+}
 - mol of Al^{3+} /volume = concentration of Al^{3+}
- Systematic errors (assumptions): "balancing" process, pH control, equilibrium reached

Experiment 2 - Volumetric Analysis: Unknown Acid Identity

- Given an unknown acid
- Goal is determine the identity of the unknown acid (monoprotic)
- Chemistry used: acid base rxn
- Method:
 - (1) standardised NaOH ($\text{NaOH} \rightarrow \text{NaCl}$)
 - NaOH concentration is tested by base with alcohol- CO_2 & benzoic acid
 - in NaCl is better than NaOH b/c it has more NaCl b/c (rough and a little bit more $\rightarrow$ more NaCl)
 - Indicator = phenolphthalein
 - (2) titration $\rightarrow$ water stain
 - (3) pH meter $\rightarrow$ pH
 - pH 8 $\rightarrow \log(\text{aq}/\text{aq})$ = after chemical point
 - H^+ is as important aq/aq = normal
- Calculations:
 - (volume of NaOH)/(molar mass of NaOH) $\rightarrow$ moles of NaOH = moles of unknown
 - (mass of unknown)/(moles of unknown) = molar mass
 - H^+ then start $\rightarrow$ greatest rate of change = eq point
- Acid dissociation constant (K_a)
 - $\text{HA} + \text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{A}^-$
 - $K_a = [\text{H}^+][\text{A}^-]/[\text{HA}]$
- Henderson-Hasselbalch equation $\text{pH}_a = \text{pH} - \log([\text{A}^-]/[\text{HA}])$
 - At 50% point: $[\text{A}^-]/[\text{HA}] = 1/1$, and $\log(1) = 0$ so $\text{pH}_a = \text{pH}$
 - At 10% point: $[\text{A}^-]/[\text{HA}] = 1/9$ and $\log(1/9) = -0.95$ so $\text{pH}_a = \text{pH} + 0.95$
 - At 90% point: $[\text{A}^-]/[\text{HA}] = 9/1$, and $\log(9) = 0.95$ so $\text{pH}_a = \text{pH} - 0.95$

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Guillaume Favre

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1. Understanding the eBook Chem 7l Ucsd Final Exam
 - The Rise of Digital Reading Chem 7l Ucsd Final Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 7l Ucsd Final Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 7l Ucsd Final Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 7l Ucsd Final Exam
 - Personalized Recommendations
 - Chem 7l Ucsd Final Exam User Reviews and Ratings
 - Chem 7l Ucsd Final Exam and Bestseller Lists
5. Accessing Chem 7l Ucsd Final Exam Free and Paid eBooks
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 - Chem 7l Ucsd Final Exam eBook Subscription Services
 - Chem 7l Ucsd Final Exam Budget-Friendly Options
6. Navigating Chem 7l Ucsd Final Exam eBook Formats

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- Chem 71 Ucsd Final Exam Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Interactive Elements Chem 71 Ucsd Final Exam
- 8. Staying Engaged with Chem 71 Ucsd Final Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 71 Ucsd Final Exam
- 9. Balancing eBooks and Physical Books Chem 71 Ucsd Final Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 71 Ucsd Final Exam
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chem 71 Ucsd Final Exam
 - Setting Reading Goals Chem 71 Ucsd Final Exam
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem 71 Ucsd Final Exam
 - Fact-Checking eBook Content of Chem 71 Ucsd Final Exam
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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