

Chemistry 1111K, Fall 2019
Problem Set 4
Problems 1 through 10 are optional
100-200 pts

Problem 1 (10 pts)

Calculate the pH of a 0.10 M Na_2S solution. $K_{\text{a1}}(\text{H}_2\text{S}) = 1.0 \times 10^{-7}$ and $K_{\text{a2}}(\text{HS}^-) = 1.0 \times 10^{-14}$.

Problem 2

Two acids, acetic acid and propanoic acid, are both weak acids. Acetic acid has a pK_{a} of 4.75 and propanoic acid has a pK_{a} of 4.87. Which acid is stronger? Explain your answer. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)

Problems 3 through 10

- From the following data, calculate the pK_{a} of H_2S and HS^- .

$\text{H}_2\text{S} + \text{H}_2\text{O} \rightleftharpoons \text{HS}^- + \text{H}_3\text{O}^+ \quad K_{\text{a1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$	$\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$
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- Calculate the pH of a 0.10 M H_2S solution. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)

$\text{H}_2\text{S} + \text{H}_2\text{O} \rightleftharpoons \text{HS}^- + \text{H}_3\text{O}^+ \quad K_{\text{a1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$	$\text{H}_2\text{S} + \text{H}_2\text{O} \rightleftharpoons \text{HS}^- + \text{H}_3\text{O}^+ \quad K_{\text{a1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$
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- Calculate the pH of a 0.10 M HS^- solution. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)

$\text{H}_2\text{S} + \text{H}_2\text{O} \rightleftharpoons \text{HS}^- + \text{H}_3\text{O}^+ \quad K_{\text{a1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$	$\text{H}_2\text{S} + \text{H}_2\text{O} \rightleftharpoons \text{HS}^- + \text{H}_3\text{O}^+ \quad K_{\text{a1}} = 1.0 \times 10^{-7}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{S}^{2-} + \text{H}_3\text{O}^+ \quad K_{\text{a2}} = 1.0 \times 10^{-14}$ $\text{HS}^- + \text{H}_2\text{O} \rightleftharpoons \text{H}_2\text{S} + \text{OH}^- \quad K_{\text{b1}} = 1.0 \times 10^{-7}$
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Problem 11

Two acids, acetic acid and propanoic acid, are both weak acids. Acetic acid has a pK_{a} of 4.75 and propanoic acid has a pK_{a} of 4.87. Which acid is stronger? Explain your answer. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)



Calculate the pH of a 0.10 M CH_3COOH solution. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)

Two acids, acetic acid and propanoic acid, are both weak acids. Acetic acid has a pK_{a} of 4.75 and propanoic acid has a pK_{a} of 4.87. Which acid is stronger? Explain your answer. (You may use the relationship $\text{pK}_{\text{a}} = -\log K_{\text{a}}$.)

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