

### Acid/Base Titration Practice Problems – Set 3

1. What is the pH of the solution when 25.00 mL 0.016 M lactic acid ( $\text{HC}_3\text{H}_5\text{O}_3$ ,  $K_a = 1.4 \times 10^{-4}$ ) is titrated with the following volumes of 0.010 M NaOH?

A. 0.00 mL

$$1.4 \times 10^{-4} = (x)(x)/(0.016 - x)$$

$$x = 0.00143$$

$$\text{pH} = -\log(0.00143) = 2.84$$

B. 10.00 mL



|   |         |         |   |     |
|---|---------|---------|---|-----|
| B | 0.00040 | 0.00010 | 0 | --- |
|---|---------|---------|---|-----|

|   |    |    |    |     |
|---|----|----|----|-----|
| C | -x | -x | +x | --- |
|---|----|----|----|-----|

|   |           |           |   |     |
|---|-----------|-----------|---|-----|
| A | 0.00040-x | 0.00010-x | x | --- |
|---|-----------|-----------|---|-----|

|  |         |   |         |     |
|--|---------|---|---------|-----|
|  | 0.00030 | 0 | 0.00010 | --- |
|--|---------|---|---------|-----|

$$\text{pH} = 3.45 \text{ (3.38... Henderson-Hasselbalch answer)}$$

C. 20.00 mL



|   |         |         |   |     |
|---|---------|---------|---|-----|
| B | 0.00040 | 0.00020 | 0 | --- |
|---|---------|---------|---|-----|

|   |    |    |    |     |
|---|----|----|----|-----|
| C | -x | -x | +x | --- |
|---|----|----|----|-----|

|   |           |           |   |     |
|---|-----------|-----------|---|-----|
| A | 0.00040-x | 0.00020-x | x | --- |
|---|-----------|-----------|---|-----|

|  |         |   |         |     |
|--|---------|---|---------|-----|
|  | 0.00020 | 0 | 0.00020 | --- |
|--|---------|---|---------|-----|

$$\text{pH} = 3.88 \text{ (3.85... Henderson-Hasselbalch answer)}$$

D. 30.00 mL



|   |         |         |   |     |
|---|---------|---------|---|-----|
| B | 0.00040 | 0.00030 | 0 | --- |
|---|---------|---------|---|-----|

|   |    |    |    |     |
|---|----|----|----|-----|
| C | -x | -x | +x | --- |
|---|----|----|----|-----|

|   |           |           |   |     |
|---|-----------|-----------|---|-----|
| A | 0.00040-x | 0.00030-x | x | --- |
|---|-----------|-----------|---|-----|

|  |         |   |         |     |
|--|---------|---|---------|-----|
|  | 0.00010 | 0 | 0.00030 | --- |
|--|---------|---|---------|-----|

$$\text{pH} = 4.35 \text{ (4.33... Henderson-Hasselbalch answer)}$$

E. 40.00 mL



|   |         |         |   |     |
|---|---------|---------|---|-----|
| B | 0.00040 | 0.00040 | 0 | --- |
|---|---------|---------|---|-----|

|   |    |    |    |     |
|---|----|----|----|-----|
| C | -x | -x | +x | --- |
|---|----|----|----|-----|

|   |           |           |   |     |
|---|-----------|-----------|---|-----|
| A | 0.00040-x | 0.00040-x | x | --- |
|---|-----------|-----------|---|-----|

|  |   |   |         |     |
|--|---|---|---------|-----|
|  | 0 | 0 | 0.00040 | --- |
|--|---|---|---------|-----|

ICE table with  $\text{C}_3\text{H}_5\text{O}_3^- + \text{H}_2\text{O} \rightleftharpoons \text{HC}_3\text{H}_5\text{O}_3 + \text{OH}^-$  ( $K_b$ )

$$7.14 \times 10^{-11} = (x)(x)/[0.00040 \text{ mol}/0.0650 \text{ L}] - x)$$

$$x = 6.63 \times 10^{-7}; \text{ so } \text{pOH} = 6.18 \text{ and } \text{pH} = 7.82$$

# Chem Fax Acid Base Titrations Answers

**Louis Meites, Henry Carrison Thomas**



## **Chem Fax Acid Base Titrations Answers:**

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contains a review of the previous year and a preview of the current year      Journal of the Society of Chemical Industry  
,1928 Lists of members for 1882 1903 issued in v 1 22 after which they were published separately wanting in v 6 and v 21

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